

Electrochemical reduction potential correlation of some insect antifeedant potent 2-phenothiazinyl chalcones

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ABSTRACT

A series containing twelve substituted styryl 2-phenothiazinyl ketones have been synthesized. The electrochemical potentials of carbonyl and vinyl groups were measured. These potentials are correlated with Hammett substituent constants, F and R parameters. From the results of statistical analyses, the effects of substituent on the group reduction potentials have been discussed. The insect antifeedant activities of these chalcones have been studied using 4th instar larvae *Achoea Janata L* with castor leaf discs.

Keywords:

2-Phenothiazinyl chalcones, Electrochemistry, Reduction potential, Correlation analysis, Insect antifeedant activities

1. INTRODUCTION

Redox potentials of organic molecules are very useful for studying electrochemical behavior such as oxidation, reduction and group transformation process. These processes are essential for conversion of one functional group into another by electron transfer. Electrochemical synthesis is another kind process for the study of properties of reactant and products[1]. Boykin and co-workers[2] have studied the dc polarographic redox potential of chalcones and they found to good correlation with sigma constants. Alston and Fry[3] investigates the reduction potential correlation with Hammett sigma *p*- constants for the chalcone system having substituents in both phenyl rings. These system performs for evaluation of Bekkum, Verkade and Wepster correlation with correlation parameters. Moraleda et al[4] have studied to found good correlation of α,β -unsaturated carbonyl compounds with half-wave potential, LUMO-HOMO energies with Hammett σ_p constants. Chalcones possess a wide range of biological activities such as antibacterial [5], antifungal[6], antiviral[7], antifeedant[8], anticancer, antimalarial[9], antituberculosis[10], antiAIDS[11], antiplosmodial[12], anti-aids[9] and antioxidant[13] activities.

These multipronged activities present in different chalcones are examined against respective strains.

The insect antifeedant activities of these chalcones have been examined against 4th instar larvae's. In agricultural field the insect antifeedant activities of the compounds are very imp for control of insects on plant leafs.

There is no report available for the correlation of redox potential and current of C=C and CO of substituted styryl 2-phenothiazinyl ketones with Hammett substituent constants, F and R parameters and insect antifeedant activities in literature in the past.

Hence the authors have taken efforts to study the electrochemical behavior and insect antifeedant activities of 2-phenothiazinyl chalcones through the correlation of redox potential with Hammett substituent constants, Field and Resonance parameters.

The insect antifeedant activities of all synthesized chalcones have been studied using 4th instar larvae *Achoea Janata L* with castor *semilooper*.

2. EXPERIMENTAL

2.1. General

All chemicals were procured from Sigma-Aldrich and Merck chemical company. The purities of these synthesized chalcones were checked with their physical constants. The physical constants and spectral data of unknown chalcones are presented in Table 1. The Infrared spectra of compounds were recorded in avatar 300 infrared spectrophotometer.

The nuclear magnetic spectra of compounds were recorded in Bruker AV 400 type nuclear magnetic spectrometer applying 400 MHz frequency for recording ¹H and 100 MHz frequency for ¹³C spectrum using CDCl₃ as solvent and TMS is an internal standard. Mass spectra of compounds were recorded in Varian 500 GC-Mass spectrometer.

The Reduction potential of these chalcones was determined in HCH Instruments inc., Electrochemical Analyzer USA using platinum wire electrode in spectral grade methanol.

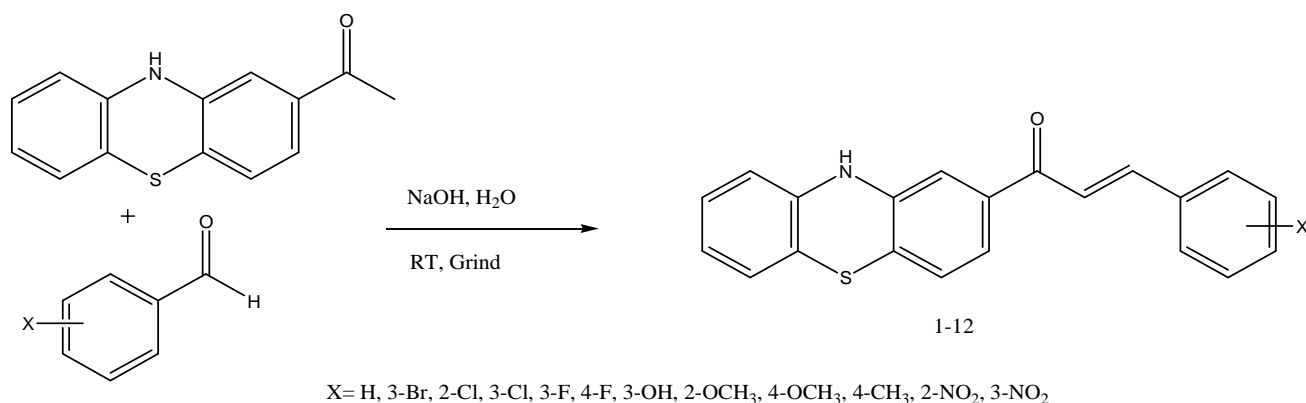
2.2. Synthesis of substituted styryl 2-phenothiazinyl ketones

All substituted styryl 2-phenothiazinyl ketones were synthesized by Green Ground Chemistry method reported in the literature[14].

Phenothiazine-2-methyl ketone (0.02mmol), substituted benzaldehydes (0.02mmol) and sodium hydroxide (0.5g) in a mortar was ground with pestle for 20 minutes (Scheme 1). Completion of the reaction was confirmed by thin layer chromatography, the reaction mixture was poured into 25ml of cold water.

The obtained brown solid was separated by filtration. The pure brown chalcones were obtained by crystallization of these brown solid with ethanol. Physical constants, analytical, UV and mass spectral data were presented in (Table 1).

Infrared spectral data of all chalcones were given in (Table 2). The NMR data of synthesized styryl 2-phenothiazinyl ketones are tabulated in (Table 3).



Scheme 1. Synthesis of 2-phenothiazinyl chalcones.

Table 1. Physical constants, analytical, Ultraviolet and mass spectral data of 2-phenothiazinyl chalcones

Entry	X	Mol. Formula	m.p. °C	M.W	UVλmax(nm)	Mass (m/z)
1	H	C ₂₁ H ₁₅ NOS	170	329	286.20	329[M ⁺], 312, 296, 286, 267, 241, 226, 198, 171, 165, 154, 148, 131, 113, 103, 85, 77, 58
2	3-Br	C ₂₁ H ₁₄ BrNOS	205	408	281.20	408[M ⁺], 410[M+2], 390, 376, 328, 310, 299, 281, 267, 241, 226, 211, 148, 182, 171, 164, 148, 127, 113, 102, 97, 73, 58
3	2-Cl	C ₂₁ H ₁₄ ClNOS	338(D.P)	363	299.00	363[M ⁺], 365[M+2], 328, 310, 296, 286, 267, 254, 241, 226, 211, 198, 183, 164, 148, 139, 113, 97, 77, 58
4	3-Cl	C ₂₁ H ₁₄ ClNOS	143	363	307.20	363[M ⁺], 365[M+2], 328, 310, 296, 286, 267, 254, 241, 226, 211, 198, 183, 164, 148, 139, 113, 97, 77, 58
5	3-F	C ₂₁ H ₁₄ FNOS	201	347	308.00	347[M ⁺], 347[M+2], 330, 314, 285, 267, 252, 226, 198, 173, 154, 149, 121, 101, 97, 73, 58
6	4-F	C ₂₁ H ₁₄ FNOS	143	347	314.60	347[M ⁺], 347[M+2], 330, 314, 285, 267, 252, 226, 198, 173, 154, 149, 121, 101, 97, 73, 58
7	3-OH	C ₂₁ H ₁₅ NO ₂ S	128	345	280.80	345[M ⁺], 327, 313, 295, 281, 267, 251, 241, 226, 221, 207, 198, 193, 169, 154, 135, 113, 99, 73, 69
8	2-OCH ₃	C ₂₂ H ₁₇ NO ₂ S	156	359	342.00	359[M ⁺], 344, 326, 316, 298, 283, 267, 254, 241, 223, 198, 183, 179, 161, 133, 118, 97, 73, 58

9	4-OCH ₃	C ₂₂ H ₁₇ NO ₂ S	162	359	341.20	359[M ⁺], 344, 326, 316, 298, 283, 267, 254, 241, 223, 198, 183, 179, 161, 133, 118, 97, 73, 58
10	4-CH ₃	C ₂₂ H ₁₇ NOS	356 (D.P.)	343	323.40	343[M ⁺], 328, 310, 300, 281, 267, 252, 241, 223, 198, 164, 145, 127, 115, 102, 91, 77, 65
11	2-NO ₂	C ₂₁ H ₁₄ N ₂ O ₃ S	324 (D.P.)	374	281.80	374[M ⁺], 358, 311, 299, 286, 267, 254, 241, 223, 198, 154, 127, 121, 102, 89, 77, 65
12	3-NO ₂	C ₂₁ H ₁₄ N ₂ O ₃ S	334 (D.P.)	374	298.00	374[M ⁺], 358, 342, 311, 299, 286, 267, 254, 241, 223, 198, 171, 154, 127, 121, 102, 89, 77, 65

Table 2. Infrared spectral data $\nu(\text{cm}^{-1})$ of 2-phenothiazinyl chalcones.

Entry	X	COs-cis	COs-trans	CH _{ip}	CH _{op}	CH=CH _{op}	>C=C< _{op}
1	H	1660.60	1627.80	1155.30	736.80	1078.10	680.80
2	3-Br	1664.50	1627.80	1155.30	740.60	1093.60	638.40
3	2-Cl	1656.70	1618.20	1178.40	752.20	1041.50	588.20
4	3-Cl	1649.00	1587.30	1191.90	734.80	1080.10	582.50
5	3-F	1651.00	1585.40	1186.10	777.30	1082.00	584.40
6	4-F	1652.90	1589.20	1184.20	736.80	1083.90	563.20
7	3-OH	1664.50	1635.50	1157.20	740.60	1093.60	638.40
8	2-OCH ₃	1652.90	1591.20	1172.60	740.60	1080.10	576.70
9	4-OCH ₃	1651.00	1587.30	1153.40	740.60	1082.00	582.50
10	4-CH ₃	1652.90	1627.80	1184.20	798.50	1083.90	565.10
11	2-NO ₂	1658.70	1598.90	1188.10	786.90	1083.90	565.10
12	3-NO ₂	1658.70	1588.30	1189.10	801.50	1084.00	588.30

Table 3. The NMR spectral chemical shift $\delta(\text{ppm})$ of 2-phenothiazinyl chalcones.

Entry	X	¹ H NMR			¹³ C NMR			
		δH_α (1H, d)	δH_β (1H, d)	Substt.	δCO	δC_α	δC_β	Substt.
1	H	7.489	7.845	---	188.79	121.89	145.16	---
2	3-Br	6.693	6.932	---	196.76	123.20	141.99	---
3	2-Cl	7.847	8.069	---	187.69	122.07	143.74	---
4	3-Cl	7.678	7.882	---	187.85	122.10	142.14	---
5	3-F	7.656	7.890	---	187.89	122.00	141.14	---
6	4-F	7.045	7.758	---	187.90	122.04	142.46	---
7	3-OH	6.681	6.926	3.336 (1H, s)	196.79	123.23	142.01	---
8	2-OCH ₃	6.708	6.955	3.819 (3H, s)	187.82	119.29	143.77	55.36 (OCH ₃)
9	4-OCH ₃	7.053	7.633	3.820 (3H, s)	187.95	122.05	142.09	55.37 (OCH ₃)
10	4-CH ₃	6.956	7.675	2.355 (3H, s)	187.92	120.76	143.78	21.03 (CH ₃)
11	2-NO ₂	7.778	7.961	---	187.67	122.78	142.16	---
12	3-NO ₂	8.108	8.874	---	188.86	122.89	142.39	---

3. RESULTS AND DISCUSSION

In the present study the ketones were chosen as an α,β -unsaturated system. These system exhibit as *E* stable conformers. The measured reduction potential (v) and the corresponding current (i) are presented in (Table 4). These potential and currents are correlated with Hammett substituent constants, F and R parameters using single and multi-regression analyses [15-17]. The results of statistical analysis of reduction potential (v) and the corresponding current (i) with Hammett substituent constants, F and R parameters are presented in (Table 5). From Table 5, correlation of reduction potential of C=C with sigma constants, F and R parameters were fail. On the other hand this current (i) were satisfactory correlated with sigma constants. The polar, resonance and fields were failing in correlation with these parameters. This is due to the inability of effect of substituents

Table 4. The reduction potential (v) and current (i) of carbonyl and C=C groups of 2-phenothiazinyl chalcones.

S. No.	Entry	Current(i) C=C	Current(i) C=O	Potential(v) C=C	Potential(v) C=O
1	H	-0.02	4.936	0.123	0.243
2	3-Br	-2.46	9.415	0.321	0.881
3	2-Cl	1.602	-2.606	1.373	-1.025
4	3-Cl	-4.088	-1.430	0.403	-0.181
5	3-F	-2.193	-4.955	0.369	-0.340
6	4-F	-1.532	-5.935	0.385	-0.097
7	3-OH	3.688	-5.376	0.365	-0.320
8	2-OCH ₃	-3.496	-4.145	0.271	-0.575
9	4-OCH ₃	-4.238	-2.246	-0.604	-0.423
10	4-CH ₃	-7.032	-1.969	-0.558	-1.314
11	2-NO ₂	-6.212	-3.531	-0.634	-1.610
12	3-NO ₂	0.283	1.954	-0.296	-3.455

not obeyed the potential and current and the conjugation between the substituent and carbonyl group shown in Figure 1. All correlation gave negative ρ values except σ_I with reduction potential of C=C and vice versa with C=C current. All correlations are fail with the reduction potential of CO with sigma constants, F and R parameters. A satisfactory correlation obtained for R parameter with reduction current of CO. Remaining sigma constants and Field parameters fail in correlations. This failure was due to conjugative structure shown in Figure 1.

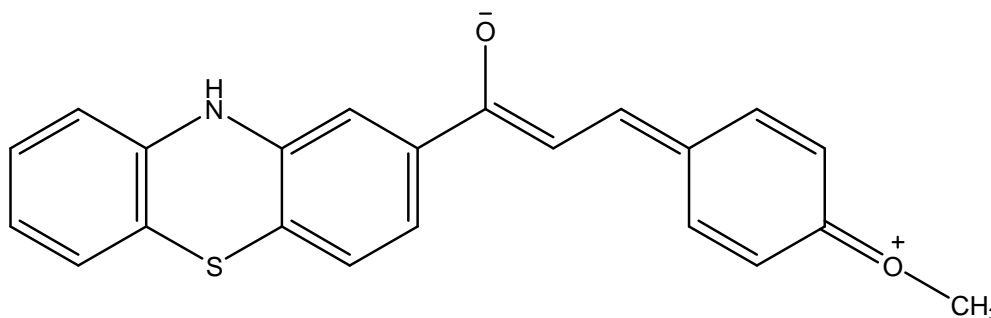


Figure 1. Resonance-conjugative structure

Table 5. Results of statistical analysis of reduction potential(v), current(i) of C=C and CO of substituted styryl 2-phenothiazenyl ketones with Hammett substituent constants σ , σ^+ , σ_I , σ_R and F and R parameters.

Frequency	Constants	r	I	ρ	s	n	Correlated derivatives
C=C (v)	σ	0.812	-0.398	-0.250	0.63	12	H, 3-Br, 2-Cl, 3-Cl, 3-F, 4-F, 3-OH, 2-OCH ₃ , 4-OCH ₃ , 4-CH ₃ , 2-NO ₂ , 3-NO ₂
	σ^+	0.811	-0.421	-0.146	0.63	12	H, 3-Br, 2-Cl, 3-Cl, 3-F, 4-F, 3-OH, 2-OCH ₃ , 4-OCH ₃ , 4-CH ₃ , 2-NO ₂ , 3-NO ₂
	σ_I	0.821	-0.292	-0.375	0.63	12	H, 3-Br, 2-Cl, 3-Cl, 3-F, 4-F, 3-OH, 2-OCH ₃ , 4-OCH ₃ , 4-CH ₃ , 2-NO ₂ , 3-NO ₂
	σ_R	0.828	-0.667	-0.837	0.61	12	H, 3-Br, 2-Cl, 3-Cl, 3-F, 4-F, 3-OH, 2-OCH ₃ , 4-OCH ₃ , 4-CH ₃ , 2-NO ₂ , 3-NO ₂
	F	0.803	-0.387	-0.098	0.64	12	H, 3-Br, 2-Cl, 3-Cl, 3-F, 4-F, 3-OH, 2-OCH ₃ , 4-OCH ₃ , 4-CH ₃ , 2-NO ₂ , 3-NO ₂
	R	0.822	-0.889	-0.414	0.62	12	H, 3-Br, 2-Cl, 3-Cl, 3-F, 4-F, 3-OH, 2-OCH ₃ , 4-OCH ₃ , 4-CH ₃ , 2-NO ₂ , 3-NO ₂
C=C (i)	σ	0.907	-0.729	-1.956	3.00	12	H, 3-Br, 2-Cl, 3-Cl, 3-F, 4-F, 3-OH, 2-OCH ₃ , 4-OCH ₃ , 4-CH ₃ , 2-NO ₂ , 3-NO ₂
	σ^+	0.801	-0.608	-2.020	3.00	12	H, 3-Br, 2-Cl, 3-Cl, 3-F, 4-F, 3-OH, 2-OCH ₃ , 4-OCH ₃ , 4-CH ₃ , 2-NO ₂ , 3-NO ₂
	σ_I	0.802	0.298	-2.142	3.01	12	H, 3-Br, 2-Cl, 3-Cl, 3-F, 4-F, 3-OH, 2-OCH ₃ , 4-OCH ₃ , 4-CH ₃ , 2-NO ₂ , 3-NO ₂
	σ_R	0.819	-2.719	-2.802	2.95	12	H, 3-Br, 2-Cl, 3-Cl, 3-F, 4-F, 3-OH, 2-OCH ₃ , 4-OCH ₃ , 4-CH ₃ , 2-NO ₂ , 3-NO ₂
	F	0.714	0.169	-2.954	4.89	12	H, 3-Br, 2-Cl, 3-Cl, 3-F, 4-F, 3-OH, 2-OCH ₃ , 4-OCH ₃ , 4-CH ₃ , 2-NO ₂ , 3-NO ₂
	R	0.813	-3.919	-1.782	3.18	12	H, 3-Br, 2-Cl, 3-Cl, 3-F, 4-F, 3-OH, 2-OCH ₃ , 4-OCH ₃ , 4-CH ₃ , 2-NO ₂ , 3-NO ₂
C=O (v)	σ	0.829	-0.577	-0.430	0.75	12	H, 3-Br, 2-Cl, 3-Cl, 3-F, 4-F, 3-OH, 2-OCH ₃ , 4-OCH ₃ , 4-CH ₃ , 2-NO ₂ , 3-NO ₂
	σ^+	0.823	-0.313	-0.516	0.76	12	H, 3-Br, 2-Cl, 3-Cl, 3-F, 4-F, 3-OH, 2-OCH ₃ , 4-OCH ₃ , 4-CH ₃ , 2-NO ₂ , 3-NO ₂
	σ_I	0.823	-0.751	-0.238	0.76	12	H, 3-Br, 2-Cl, 3-Cl, 3-F, 4-F, 3-OH, 2-OCH ₃ , 4-OCH ₃ , 4-CH ₃ , 2-NO ₂ , 3-NO ₂
	σ_R	0.839	-1.293	-0.838	0.72	12	H, 3-Br, 2-Cl, 3-Cl, 3-F, 4-F, 3-OH, 2-OCH ₃ , 4-OCH ₃ , 4-CH ₃ , 2-NO ₂ , 3-NO ₂
	F	0.813	-0.388	-0.369	0.77	12	H, 3-Br, 2-Cl, 3-Cl, 3-F, 4-F, 3-OH, 2-OCH ₃ , 4-OCH ₃ , 4-CH ₃ , 2-NO ₂ , 3-NO ₂
	R	0.833	-0.636	-0.750	0.74	12	H, 3-Br, 2-Cl, 3-Cl, 3-F, 4-F, 3-OH, 2-OCH ₃ , 4-OCH ₃ , 4-CH ₃ , 2-NO ₂ , 3-NO ₂
C=O (i)	σ	0.808	0.158	0.251	0.57	12	H, 3-Br, 2-Cl, 3-Cl, 3-F, 4-F, 3-OH, 2-OCH ₃ , 4-OCH ₃ , 4-CH ₃ , 2-NO ₂ , 3-NO ₂
	σ^+	0.821	0.316	0.260	0.55	12	H, 3-Br, 2-Cl, 3-Cl, 3-F, 4-F, 3-OH, 2-OCH ₃ , 4-OCH ₃ , 4-CH ₃ , 2-NO ₂ , 3-NO ₂
	σ_I	0.829	0.818	-0.021	0.55	12	H, 3-Br, 2-Cl, 3-Cl, 3-F, 4-F, 3-OH, 2-OCH ₃ , 4-OCH ₃ , 4-CH ₃ , 2-NO ₂ , 3-NO ₂
	σ_R	0.718	-0.495	0.125	0.56	12	H, 3-Br, 2-Cl, 3-Cl, 3-F, 4-F, 3-OH, 2-OCH ₃ , 4-OCH ₃ , 4-CH ₃ , 2-NO ₂ , 3-NO ₂
	F	0.822	0.859	0.051	0.56	12	H, 3-Br, 2-Cl, 3-Cl, 3-F, 4-F, 3-OH, 2-OCH ₃ , 4-OCH ₃ , 4-CH ₃ , 2-NO ₂ , 3-NO ₂
	R	0.911	-0.192	0.192	0.57	12	H, 3-Br, 2-Cl, 3-Cl, 3-F, 4-F, 3-OH, 2-OCH ₃ , 4-OCH ₃ , 4-CH ₃ , 2-NO ₂ , 3-NO ₂

r = Correlation coefficient; ρ = Slope; I = Intercept; s = Standard deviation; n = Number of substituents.

Some of the single parameter electrochemical correlations were fails with Hammett sigma constants, F and R parameters. On the other hand the multi-regressions were worthwhile and leads to satisfactory correlated with Swain-Lupton[18] F and R parameters collectively. The correlation equations 1-8 are as

$$C=C(v) = -0.540(\pm 0.042) - 0.349(\pm 0.081)\sigma_I - 0.827(\pm 0.081)\sigma_R \quad \dots(1)$$

$$(R = 0.930, n = 12, P > 90\%)$$

$$C=C(v) = 0.534(\pm 0.039) - 0.142(\pm 0.078)F - 0.420(\pm 0.050)R \quad \dots(2)$$

$$(R = 0.923, n = 12, P > 90\%)$$

$$C=C(i) = 1.622(\pm 0.291) - 3.546(\pm 0.593)\sigma_I + 11.362(\pm 5.682)\sigma_R \quad \dots(3)$$

$$(R = 0.951, n = 12, P > 95\%)$$

$$C=C(i) = 1.282(\pm 0.267) - 4.413(\pm 0.051)F + 6.271(\pm 0.141)R \quad \dots(4)$$

$$(R = 0.950, n = 12, P > 95\%)$$

$$C=O(v) = -0.565(\pm 0.049) - 0.683(\pm 0.096)\sigma_I - 1.243(\pm 0.097)\sigma_R \quad \dots(5)$$

$$(R = 0.944, n = 12, P > 90\%)$$

$$C=O(v) = -0.588(\pm 0.048) - 0.387(\pm 0.090)F - 0.636(\pm 0.059)R \quad \dots(6)$$

$$(R = 0.935, n = 12, P > 90\%)$$

$$C=O(i) = -0.177(\pm 0.370) + 0.834(\pm 0.076)\sigma_I - 0.591(\pm 0.071)\sigma_R \quad \dots(7)$$

$$(R = 0.935, n = 12, P > 90\%)$$

$$C=O(i) = -0.009(\pm 0.003) + 0.542(\pm 0.075)F - 0.170(\pm 0.045)R \quad \dots(8)$$

$$(R = 0.958, n = 12, P > 95\%)$$

3.1. Insect antifeedant activity

The multipronged activities present in different epoxy ketones are intended to examine their insect antifeedant activities against castor *semilooper*. The larvae of *Achoea Janata* L were reared as described on the leaves of castor *Ricinus communis* in the laboratory at the temperature range of $26^\circ\text{C} \pm 1^\circ\text{C}$ and a relative humidity of 75-85%. The leaf – disc bioassay method[8,19] was used against the 4th instar larvae to measure the antifeedant activity. The 4th instar larvae were selected for testing because the larvae at this stage feed very voraciously.

3.2. Measurement of insect antifeedant activity of chalcones

Leaf discs of a diameter of 1.85 cm were punched from castor leaves with the petioles intact. All chalcones. Insect antifeedant activities of substituted styryl 2-phenothiazinyl ketones were dissolved in acetone at a concentration of 200 ppm dipped for 5 minutes. The leaf discs were air-dried and placed in one litre beaker containing little water in order to facilitate translocation of water. Therefore the leaf discs remains fresh throughout the duration of the rest, 4th instar larvae of the test insect, which had been preserved on the leaf discs of all chalcones and allowed to feed on them for 24 hours. The areas of the leaf discs consumed were measured by Dethlers[19] method. The observed antifeedant activity of chalcones was presented in (Table 6).

Table 6. Insect antifeedant activities of substituted styryl 2-phenothiazenyl ketones.

Entry	X	4-6 pm	6-8 pm	8-10 pm	10-12 pm	12-6 am	6-8 am	8 am-12 Nn	12 Nn-2 pm	2-4 pm	Total leaf disc consumed in 24 hrs
1	H	1	1	0.5	0.5	0.5	1	1	1	1	8
2	3-Br	1	2	0.25	1	0.1	0	1	1	0.5	5
3	3-Cl	0.5	0.25	0.25	0.5	0.5	0.5	1	1	0.5	5
4	4-Cl	0.5	0.5	0.25	1	0.5	0.5	0.25	0.25	0.25	4
5	3-F	0.5	0.5	0.25	1	0.5	0.5	0.25	0.25	0.25	4
6	4-F	0.5	0.5	0.25	1	0.5	0	0	0	0.25	3
7	3-OH	1	2	2	1	0	0	1	1	1	9
8	2-OCH ₃	1	1	1	0.5	0.5	1	2	1	1	10
9	4-OCH ₃	1	0.5	0.5	1	1	0	1	1	1	9
10	4-CH ₃	0.5	1	0.5	2	2	0.5	0.5	1	1	7
11	3-NO ₂	2	3	3	1	1	1	0.5	1	0	12
12	4-NO ₂	2	3	3	1	0	0	0.5	1	0	7

Number of leaf discs consumed by the insect (Values are mean + SE of five).

Table 7. Insect antifeedant activity of compound **6** 4-fluorostyryl styryl 2-phenothiazenyl ketone at the concentration of 50, 100 and 150 ppm.

ppm	4-6 pm	6-8 pm	8-10 pm	10-12 pm	12-6 am	6-8 am	8am-12 Nn	12 Nn-2 pm	2-4 pm	Total leaf disc consumed in 24 hrs
50	0.5	0.5	0	0	0	0	0	0	0	0.1
100	0	0.25	0	0	0	0	0	0	0	0.05
150	0	0.5	.25	0	0.25	0	0	0	0	0.1

Number of leaf discs consumed by the insect (Values are mean + SE of five).

The results of the antifeedant activity of keto chalcones presented in (Table 7) and it reveals that the compounds **4-6** are found to reflect remarkable antifeedant among all other chalcones. This test is performed with the insects which took only two-leaf disc soaked under the solution of this compound[20].

Compounds **5** showed enough antifeedant activity but their activities are lesser than **4**. Further, compound **4** was subjected to measure the antifeedant activity at different 50, 100, 150 ppm concentrations and the observation reveals that as the concentrations are decreases, the activity also decreases.

From the results in Table 8, it is observed that the chalcone **4** showed an appreciable antifeedant activity at 150 ppm concentration.

4. CONCLUSION

There are twelve substituted styryl 2-phenothiazanyl ketones have been synthesized by Ground green chemistry method. These chalcones were characterized by their physical constants and spectral data. The electrochemical reduction potential and currents of carbonyl and vinyl parts of these chalcones have been measured. These components were correlated with Hammett substituent constants, F and R parameters. The reduction current of vinyl part and carbonyl carbons were satisfactorily correlated. The multi-regression analysis of these components correlated satisfactorily with Swain-Lupton's, F and R parameters. Insect antifeedant activities of these chalcones have been studied using 4th instar larvae *Achoea Janata L* with castor leaf discs.

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