

Received August 12, 2013; reviewed; accepted October 2, 2013

Pr-DOPED TiO₂. THE EFFECT OF METAL CONTENT ON PHOTOCATALYTIC ACTIVITY

Joanna RESZCZYŃSKA^{*}, Daniel ARENAS ESTEBAN^{,**}, Maria GAZDA^{***},
Adriana ZALESKA^{*}**

^{*} Department of Chemical Technology, Faculty of Chemistry, Gdansk University of Technology, Gdansk, Poland, e-mail: adriana.zaleska@pg.gda.pl

^{**} Faculty of Chemical Sciences, Complutense University of Madrid, Madrid, Spain

^{***} Department of Solid State Physics, Faculty of Applied Physics and Mathematics, Gdansk University of Technology, 80-952 Gdansk, Poland

Abstract: Pr-TiO₂ nanoparticles were prepared by using a sol-gel method. As-prepared samples were characterized by BET measurements, X-ray powder diffraction analysis (XRD) and UV-Vis spectra. Visible and ultraviolet light photocatalytic activity of the sample was studied by photodegradation of phenol, while considering the influence of the dopant concentration. TiO₂ doped with 0.25 mol% of praseodymium showed the highest photocatalytic activity under visible light irradiation. Pr-TiO₂ had anatase structure. The surface area was higher for powders with higher content of rare earth metal ion, and ranged from 121 to 150 m²/g. Red shifts of absorption edge toward the visible region were observed for the doped samples compared to pure TiO₂.

Keywords: photocatalysis, rare earth metal, praseodymium, modified TiO₂, phenol

Introduction

Phenols are generally considered to be one of important organic pollutants discharged into the environment causing unpleasant taste and odor of drinking water (Grabowska et al., 2012a). Phenol is toxic even at low concentrations, and also its presence in natural waters can lead further to formation of substituted compounds during the disinfection and oxidation processes (Busca et al., 2008). In response, it has become a challenge to achieve an effective removal of persistent organic pollutants from wastewater effluent to minimize the risk of pollution problems from such toxic chemicals and to enable its reuse. TiO₂ is a commonly used photocatalyst in water purification, wastewater and gas streams treatment (Gorska et al., 2008; Zielinska-Jurek et al., 2010; Grabowska et al., 2012b; Nischk et al., 2013). However, utilization

of pure TiO₂ is significantly constrained because of its wide band gap, low efficiency of sunlight utilization, low quantum efficiency, and high recombination rate of electrons-holes. Therefore, it is of great interest to improve the photocatalytic activity of TiO₂ in order to enlarge practical applications of TiO₂-based photocatalysts (Shi et al., 2012). Rare-earth ion doped materials have played an important role in the development of optical communication technology during the past few decades (Shang et al., 2008). Lanthanide ions are known for their ability to form complexes with various Lewis bases (e.g., acids, amines, aldehydes, alcohols, thiols, etc.) in interaction of these functional groups with the f-orbitals of lanthanides (Ranjit et al., 2001). Titania doping with lanthanides is a method to shift the maximum of absorption as well as enhance the photocatalytic activity (Hassan et al., 2012). Rare earth metals could be introduced to the surface of TiO₂ by various methods such as: sol-gel (Sheng et al., 2011; Tang et al., 2012), hydrothermal (Pedroni et al., 2012; Li et al., 2013), electrospinning (Cacciotti et al., 2011; Hassan et al., 2012) and surface impregnation (Parida and Sahu, 2008).

In our previous paper (Reszczyńska et al., 2012) the sol-gel method to synthesize Er and Yb modified TiO₂ nanoparticles was presented. It was observed that the dopant amount affected photocatalytic activity of semiconductor. The presence of erbium was found more beneficial for visible light activation of TiO₂ doped photocatalysts than ytterbium.

In this work, a series of Pr-doped TiO₂ nanoparticles were synthesized using the sol-gel technique with various Pr concentrations. The photocatalytic activity of as-prepared Pr-doped TiO₂ was evaluated using degradation of phenol in aqueous solutions under UV and Vis irradiation. The BET surface area, X-ray powder diffraction analysis (XRD) and UV-Vis spectra were evaluated.

Experimental

Materials and instruments

Titanium(IV) isopropoxide (97%) was purchased from Sigma-Aldrich Co. and used as titanium source for the preparation of TiO₂ nanoparticles. Pr(NO₃)₃·5H₂O (99.99%), from Sigma–Aldrich Co. was used as rare earth metal source in the preparation procedure. Acetic acid and ethanol (POCH S.A. Poland) were used without further purification. Deionized water was used for all reactions and treatment processes. A commercial form of TiO₂ (P25, crystalline composition: 80% anatase, 20% rutile) from Evonik, Germany was used for the comparison of the photocatalytic activity.

The nitrogen adsorption-desorption isotherms were recorded at liquid nitrogen temperature (77 K) on a Micromeritics Gemini V (model 2365) and the specific surface areas were determined by the Brunauer-Emmett-Teller (BET) method in a relative pressure (p/p_0) range of 0.05–0.3 Pa. All the samples were degassed at 200 °C prior to nitrogen adsorption measurement.

DRS UV–Vis spectra of synthesized materials were recorded in a scan range 350–700 nm using a UV–Vis spectrophotometer (UV-Vis Thermo model: Nicolet Evolution 220 with ISA-220 integrating sphere) equipped with an integrating sphere and BaSO₄ was used as the reference. X-ray diffraction patterns (XRD) were recorded on a X-ray diffractometer (Xpert PRO-MPD, Philips) with Cu target K_α-ray ($\lambda=0.15404$ nm) in the range $2\theta=20-68^\circ$. The patterns were analyzed by the Rietveld refinement method using LHPM program (Hill and Howard, 1986). The crystallite size was estimated by Scherrer equation. The accuracy of the grain size analysis was estimated to be about 20%.

PREPARATION OF Pr DOPED TiO₂ PHOTOCATALYSTS

For the preparation of photocatalyst the sol-gel method was used. TiO₂-based photocatalysts were obtained according to the procedures presented by a simplified block diagram in Fig. 1. The sample of 15 cm³ titanium isopropoxide was added to the mixture of 3.0 cm³ acetic acid and 60 cm³ ethanol. The sol was stirred for 1 h, then 20 cm³ of Pr(NO₃)₃ water solution, with pH adjusted to 3.5 (with 65 % HNO₃), was added dropwise. The obtained solution was stirred for 2.5 h. The prepared gel was filtered and washed four times with water using a centrifugal and dried at 80 °C. The heat-treatment of 400 °C for 2.5 h in the air was applied to the sample (heating rate of 5 °C/min). The concentration of metal precursors, which varied from 0 to 1.5 mol%, was related to the concentration of TIP in the sol-gel system.

MEASUREMENTS OF PHOTOCATALYTIC ACTIVITY

The photocatalytic activity of Pr-TiO₂ powders in the visible and ultraviolet lights were estimated by measuring the decomposition rate of phenol in an aqueous solution. A 25 cm³ of photocatalyst suspension (125 mg) in 0.21 mM phenol aqueous solution was placed in a quartz photoreactor. After 30 min aeration (5 dm³/h) the suspension was irradiated with a Xenon 1000 W lamp. The optical path included water filter to cut off IR irradiation. For the test of visible-light-induced activity the light beam was passed through a GG420 filter to cut-off wavelengths shorter than 420 nm. One milliliter aliquots of aqueous suspension were collected at regular time periods during irradiation and filtered through the syringe filters ($\varnothing=0.2$ mm) to remove the photocatalyst particles. The temperature of suspension was maintained at 10 °C using a thermostatically controlled water bath. Phenol concentration was estimated by the colorimetric method ($\lambda=480$ nm) after derivatisation with diazo-pnitroaniline using the UV-Vis spectrophotometer (DU-520, Beckman).

Results and discussion

The sample labels, BET surface areas, crystal size and photoactivity for praseodymium-doped TiO₂ photocatalysts were listed in Table 1. For all the doped samples prepared by the sol-gel method, the BET surface area was higher than for the

commercially available P25 ($50 \text{ m}^2/\text{g}$). The surface area was higher for powders with higher content of rare earth metal ion, and ranged from 121 to $150 \text{ m}^2/\text{g}$. The Pr-doped sample has larger BET surface area than the undoped sample, indicating that the Pr-doping increases the thermal stability of mesoporous framework. Based on the literature data, it is likely that lanthanides ions are not incorporated into the matrix of TiO_2 but rather form rare-earth oxides, RE_2O_3 (Yang et al., 2011).

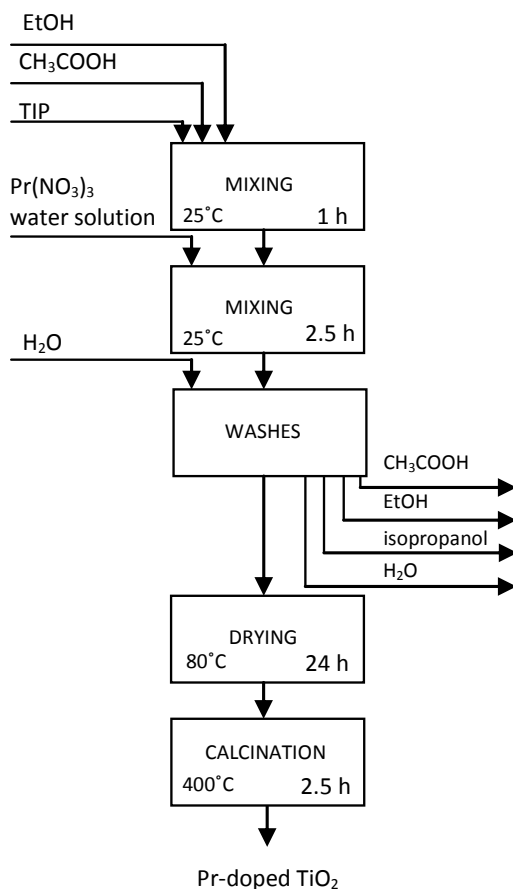


Fig. 1. Block diagram of Pr-doped TiO_2 prepared by the sol-gel method

The melting point of Pr_2O_3 is 2138°C is higher than that of TiO_2 . The surface coverage of metal oxide with higher melting point probably hampers the adjustment of matrix particles and rearrangement of crystal. Consequently, the Pr-doping improves the microstructural properties of TiO_2 . Similar observation was made for Yb/N- TiO_2 photocatalysts (Yan et al., 2013). Rare earth ions doping can reduce the crystallite size but increase the surface area of TiO_2 , which also contribute to the enhancement of

adsorption capacity of TiO₂ for organic pollutants (Liang et al., 2006; 2009). The exceeded surface area for lanthanoid doped TiO₂, in comparison with the pure one, was also reported by Bingham and Daoud (2011).

Table 1. Surface characteristics and photocatalytic activity of Pr-TiO₂ photocatalysts

Photocatalysts type	Amount of metal precursor (mol%)	S _{BET} (m ² /g)	Average crystallite size (nm)	Phenol degradation rate under Vis irradiation (λ > 420 nm) (μmol/dm ³ ·min)	Phenol degradation rate under UV-Vis irradiation (μmol/dm ³ ·min)
TiO ₂ _Pure	none	130	11.1	0.42	2.80
TiO ₂ _Pr(0.1)	0.1	121	11.2	0.69	2.77
TiO ₂ _Pr(0.25)	0.25	146	11.0	1.01	2.55
TiO ₂ _Pr(0.5)	0.5	135	10.0	0.77	2.84
TiO ₂ _Pr(0.7)	0.7	139	9.6	0.56	3.17
TiO ₂ _Pr(1.0)	1.0	148	9.7	0.54	1.48
TiO ₂ _Pr(1.5)	1.5	150	9.6	0.75	3.12
P25	none	50	not measured	0.34	3.19

To study the optical absorption properties of as-prepared samples, the UV-Vis absorption spectra in the range 350–700 nm were investigated, and the results are shown in Fig. 2. It can be seen that modification of titania with praseodymium significantly affected the light absorption property of photocatalysts. Red shifts of

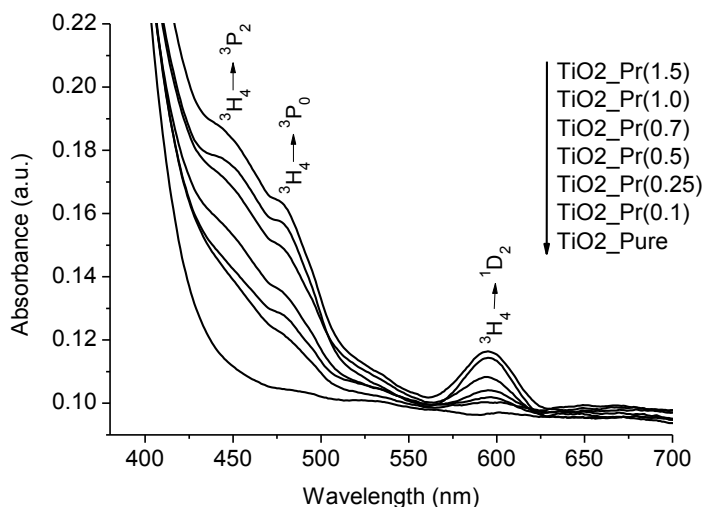


Fig. 2. Absorption properties of Pr-TiO₂ prepared by the sol-gel method

absorption edge toward the visible region were observed for the doped samples compared to pure TiO_2 . The correlation between the amount of rare earth metal and the absorption edge shift to a longer wavelength was observed. Furthermore, it can be seen that there are three typical for praseodymium absorption peaks located at 446, 480 and 595 nm and they are attributed to transitions of $^3\text{H}_4$ to $^3\text{P}_2$, $^3\text{P}_0$ and $^1\text{D}_2$ (Su et al., 2008; Yang et al., 2011). The intensity of rare earth ion absorption bands was found to increase with increasing in the rare earth ion content.

The X-ray diffraction analysis (XRD) was used to verify the crystalline phase of synthesized doped TiO_2 and to estimate the crystallite size. The XRD patterns of examined samples are shown in Fig. 3. In all prepared photocatalysts, Fig. 3 presents a group of lines at 2θ values of 25.3° , 37.8° , 48.0° , 53.9° and 55.1° , which are attributed to anatase phase (El-bahy et al., 2009). However, no peaks related to the presence of praseodymium or praseodymium oxides were detected for Pr- TiO_2 photocatalysts. Thus, it indicates that the dopant nanoparticles at the TiO_2 surface are very small and highly dispersed (Zhou and He, 2012) or the concentration of rare earth metal was too low for the XRD to reveal (Wu et al., 2011). In our data the average particle size, determined on the base of Scherrer equations, ranged from 9.6 to 11.2 nm for $\text{TiO}_2\text{-Pr}(1.5)$ and $\text{TiO}_2\text{-Pr}(0.1)$, respectively. For undoped TiO_2 the particle size was 11.2 nm (Table 1). The particle size was found to decrease with increase in metal ion dopant concentration in TiO_2 .

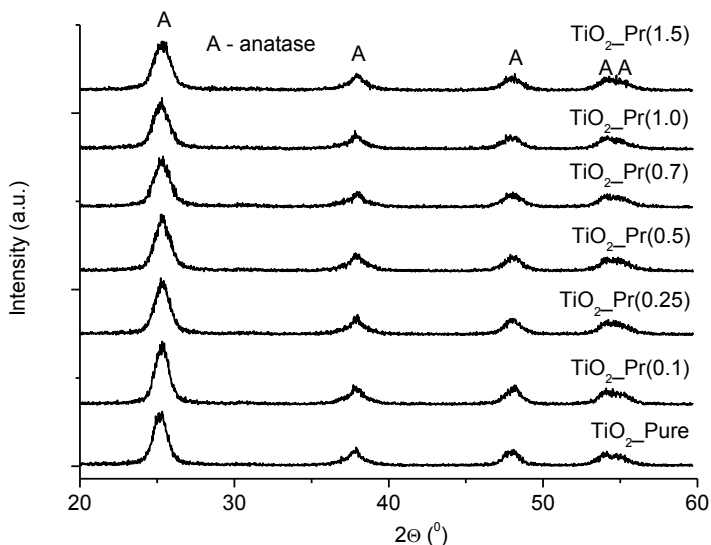


Fig. 3. XRD pattern of Pr- TiO_2 photocatalysts prepared by sol-gel method

The photocatalytic activity of obtained TiO_2 powders was estimated by measuring the decomposition rate of phenol in the aqueous solution in the presence of UV-Vis or

visible light irradiation ($\lambda > 420$ nm). No degradation of phenol was observed in the absence of photocatalysts or illumination. The pure TiO₂ synthesized by the sol-gel method without any dopant and P25 were used as the reference system. Kinetics of phenol photodegradation in aqueous solution under Vis/UV-Vis light irradiation in the presence of Pr-doped TiO₂ nanoparticles is shown in Figs. 4a,b, respectively. The observed rate constants are listed in Table 1.

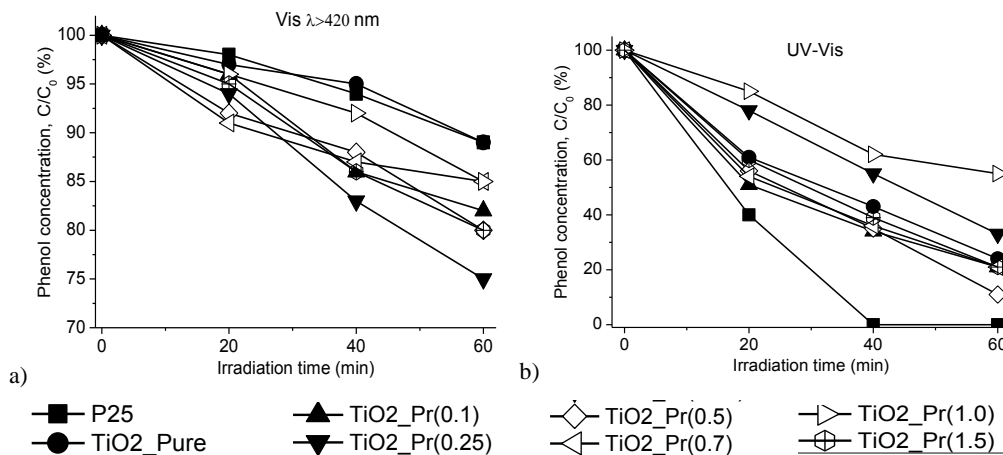


Fig. 4. Efficiency of phenol photodegradation in the presence of Pr-doped TiO₂ during irradiation light: a) Vis ($\lambda > 420$ nm), b) UV/Vis

Phenol degradation under visible light irradiation with the participation of praseodymium doped TiO₂ is presented in Fig. 4a. The highest photoactivity was observed for the sample prepared with 0.25 mol% Pr. After 60 min irradiation of aqueous suspension initially containing 20 mg/dm³ of phenol, 25 % of those was degraded. All rare earth metal-doped photocatalysts had better photocatalytic activity than the undoped ones and P25 under visible light irradiation. The phenol degradation rate under UV-Vis irradiation increased from 2.55 to 3.17 $\mu\text{mol}/\text{dm}^3 \cdot \text{min}$ for 0.25 to 0.7 mol% praseodymium-doped TiO₂, respectively. Praseodymium and nitrogen co-doped titania (Pr/N-TiO₂) photocatalysts, which could degrade Bisphenol A (BPA) under visible light irradiation, were prepared by the modified sol-gel process by Yang et al. (2011). It was found that Pr doping inhibited the growth of crystalline size and transformation from anatase to rutile. For the best photocatalytic activity, the optimal dopant amount of Pr was 1.2 mol% and the calcination temperature was 500 °C. Pr doping could slow the radiative recombination of photogenerated electrons and holes in TiO₂. The improvement of photocatalytic activity was ascribed to the synergistic effects of nitrogen and Pr co-doping.

Chiou and Juang (2007) prepared the Pr-doped TiO₂ nanoparticles with a composition of Ti_{1-x}Pr_xO₂ (x = 0.018–0.22) using the acid-peptized sol-gel method. The Pr-doped TiO₂ showed high activity for photocatalytic degradation of phenol.

Light absorption reached the maximum for the particles with 0.072 mol% Pr, which was consistent with the trends of phenol photodegradation efficiency.

It is presumed that rare earth doping can lead to defects in the TiO₂ lattice and these defects act as electron traps. The doping may also favor separation of charge carriers efficiently and enhance the photocatalytic activity. It is suggested that rare earth oxides are located on the surface of titania, but more detailed studies about the mechanism of its photoexcitation is required (Wang et al., 2010).

Conclusion

In summary, the Pr-TiO₂ nanoparticles were prepared by the sol-gel method and had the anatase structure. Photocatalysts based on TiO₂ and doped with praseodymium caused an increase of photocatalytic activity under visible light irradiation. It was observed that the dopant affects the surface area and crystal size of TiO₂ powder samples. The photocatalytic activity depends on rare earth ion precursor concentration used during preparation. TiO₂ doped with 0.25 mol% of praseodymium showed the highest photocatalytic activity under visible light. After 60 min of irradiation, 25 % of phenol was degraded. The absorption spectra of Pr-TiO₂ samples show stronger absorption in the UV-Vis region than pure TiO₂.

Acknowledgements

This research was financially supported by Polish National Science Center (grant No. 2011/01/N/ST5/05537).

References

- BINGHAM S., DAOUD W., 2011, *Recent advances in making nano-sized TiO₂ visible-light active through rare-earth metal doping*, J. Mater. Chem. 21, 2041–2050.
- BUSCA G., BERARDINELLI S., RESINI C., ARRIGHI, L., 2008, *Technologies for the removal of phenol from fluid streams: A short review of recent developments*, J. Hazard. Mater. 160, 265–288.
- CACCIOTTI I., BIANCO A., PEZZOTTI G., GUSMANO G., 2011, *Terbium and ytterbium-doped titania luminescent nanofibers by means of electrospinning technique*, Mater. Chem. Phys. 126, 532–541.
- CHIOU C.-H., JUANG R.-S., 2007, *Photocatalytic degradation of phenol in aqueous solutions by Pr-doped TiO₂ nanoparticles*, J. Hazard. Mater. 149, 1–7.
- GORSKA P., ZALESKA A., KOWALSKA E., KLIMCZUK T., SOBCZAK J. W., SKWAREK E., JANUSZ, W., HUPKA, J., 2008, *TiO₂ photoactivity in Vis and UV light: the influence of calcination temperature and surface properties*, Appl. Catal., B 84, 440–447.
- GRABOWSKA E., RESZCZYNSKA J., ZALESKA A., 2012a, *Mechanism of phenol photodegradation in the presence of pure and modified-TiO₂: A review*, Water Res. 46(17), 5453–5471.
- GRABOWSKA E., SOBCZAK J., GAZDA M., ZALESKA A., 2012b, *Surface properties and visible light activity of W-TiO₂ photocatalysts prepared by surface impregnation and sol-gel method*, Appl. Catal., B 117–118, 351–359.
- HASSAN M. S., AMNA T., YANG O.-B., KIM H.-C., KHIL M.-S., 2012, *TiO₂ nanofibers doped with rare earth elements and their photocatalytic activity*, Ceram. Int. 38, 5925–5930.

- HILL R. J., HOWARD C. J., 1986, *The High Score plus Rietveld algorithm is based on the source codes of the program LHPM*, A computer program for Rietveld analysis of fixed wavelength X-ray and neutron diffraction patterns, Australian Atomic Energy Commission Research Report M112.
- LI H., ZHENG K., SHENG Y., SONG Y., ZHANG H., HUANG J., HUO Q., ZOU H., 2013, *Facile synthesis and luminescence properties of TiO₂:Eu³⁺ nanobelts*, Opt. Laser Technol. 49, 33–37.
- LIANG C.-H., HOU M.-F., ZHOU S.-G., LI F.-B., LIU C.-S., LIU T.-X., GAO Y.-X., WANG X.-G., LÜ, J.-L., 2006, *The effect of erbium on the adsorption and photodegradation of orange I in aqueous Er³⁺-TiO₂ suspension.*, J. Hazard. Mater. 138, 471478.
- LIANG C.-H., LIU C.-S., LI F.-B., WU F., 2009, *The effect of praseodymium on the adsorption and photocatalytic degradation of azo dye in aqueous Pr³⁺-TiO₂ suspension*, Chem. Eng. J. 147, 219–225.
- NISCHK M., MAZIERSKI P., GAZDA M., ZALESKA A., 2013, *Ordered TiO₂ nanotubes: The effect of preparation parameters on the photocatalytic activity in air purification process*, Appl. Catal., B 144, 674–685.
- PARIDA K. M., SAHU N., 2008, *Visible light induced photocatalytic activity of rare earth titania nanocomposites*, J. Mol. Catal. A: Chem. 287, 151-158.
- PEDRONI M., PICCINELLI F., POLIZZI S., SPEGHINI A., BETTINELLI M., HARO-GONZÁLEZ P., 2012, *Upconverting Ho-Yb doped titanate nanotubes*, Mater. Lett. 80, 81-83.
- RANJIT K. T., WILLNER I., BOSSMANN S. H., BRAUN A. M., 2001, *Lanthanide oxide-doped titanium dioxide photocatalysts: Novel photocatalysts for the enhanced degradation of p-chlorophenoxyacetic acid*, Environ. Sci. Technol. 35, 1544-1549.
- RESZCZYNSKA J., IWULSKA A., ŚLIWINSKI G., ZALESKA A., 2012, *Characterization and photocatalytic activity of rare earth metal-doped titanium dioxide* Physicochem. Probl. Miner. Process. 48(1), 201-208.
- SHANG Q., YU H., KONG X., WANG H., WANG X., SUN Y., ZHANG Y., ZENG Q., 2008, *Green and red up-conversion emissions of Er³⁺-Yb³⁺ co-doped TiO₂ nanocrystals prepared by sol-gel method*, J. Lumin. 128, 1211-1216.
- SHENG Y., ZHANG L., LI H., XUE J., ZHENG K., GUO N., HUO Q., ZOU H., 2011, *Photoluminescence of TiO₂ films co-doped with Tb³⁺/Gd³⁺ and energy transfer from TiO₂/Gd³⁺ to Tb³⁺ ions*, Thin Solid Films 519, 7966-7970.
- SHI H., ZHANG T., AN T., LI B., WANG X., 2012, *Enhancement of photocatalytic activity of nano-scale TiO₂ particles co-doped by rare earth elements and heteropolyacids*, J. Colloid Interface Sci. 380, 121-127.
- SU W., CHEN J., WU L., WANG X., WANG X., FU X., 2008, *Visible light photocatalysis on praseodymium(III)-nitrate-modified TiO₂ prepared by an ultrasound method*, Appl. Catal., B 77, 264-271.
- TANG J., CHEN X., LIU Y., GONG W., PENG Z., CAI T., JIN L., DENG Q., 2012, *Europium-doped mesoporous anatase with enhanced photocatalytic activity toward elimination of gaseous methanol*, J. Phys. Chem. Solids 73, 198-203.
- WANG C., YANHUI A., WANG P., HOU J., QIAN J., 2010, *Preparation, characterization and photocatalytic activity of the neodymium-doped TiO₂ hollow spheres*, Appl. Surf. Sci. 254, 227-231.
- WU J., LIU Q., GAO P., ZHU Z., 2011, *Influence of praseodymium and nitrogen co-doping on the photocatalytic activity of TiO₂*, Mater. Res. Bull. 46, 1997-2003.
- YAN P., JIANG H., ZANG S., LI J., WANG Q., WANG Q., 2013, *Sol-solvothermal preparation and characterization of (Yb, N)-codoped anatase-TiO₂ nano-photocatalyst with high visible light activity*, Mater. Chem. Phys. 139, 1014-1022.
- YANG J., DAI J., LI J., 2011, *Synthesis, characterization and degradation of Bisphenol A using Pr, N co-doped TiO₂ with highly visible light activity*, Appl. Surf. Sci. 257, 8965–8973.

- ZHOU W., HE Y., 2012, *Ho/TiO₂ nanowires heterogeneous catalust with enhanced photocatalytic properties by hydrothermal synthesis method*, Chem. Eng. J. 179, 412–416.
- ZIELINSKA-JUREK A., WALICKA M., TADAJEWSKA A., LACKA I., GAZDA M., ZALESKA A., 2010, *Preparation of Ag/Cu-doped titanium (IV) oxide nanoparticles in w/o microemulsion*, Physicochem. Probl. Miner. Process. 45, 113–126.