

Synthesis and spectral studies of some *N*-[(*E*)-phenylmethylidene]benzenesulfonamides

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ABSTRACT

A series of some *N*-[(*E*)-phenylmethylidene]benzenesulfonamide derivatives have been synthesised using solid SiO₂-H₃PO₄ catalyst under solvent free conditions in microwave irradiation. The synthesised *E*-*N*-benzene sulfonilimines purities have been verified by their physical constants and spectroscopic data. The spectral frequencies are correlated with Hammett substituent constants, F and R parameters using linear regression analysis. From the results of statistical analysis the effect of substituents on the group frequencies will be discussed.

Keywords: *N*-[(*E*)-phenylmethylidene]benzenesulfonamides; SiO₂-H₃PO₄; Greener synthesis; Hammett correlations.

1. INTRODUCTION

The *E*-*N*-benzene sulfonilimines are versatile useful intermediates for organic synthesis [1]. These imines possess the -CH=N-SO₂- moiety and this is bonded with alkyl - aryl or aryl- aryl groups in both sides. *N*-tosylimine intermediates are used as excellent substrates in Diels-Alder reactions and it is called as Aza-Diels-Alder reaction [2]. In this Diels-Alder reactions these intermediates undergo nucleophilic additions [3, 4], reductions [5], radical [6] and Baylis-Hillman [7, 8]. Generally, these *E*-*N*-benzene sulfonilimines were prepared by condensation of aldehydes and *N*-tosylamines using several methods such as conventional solvent assisted condensation [9, 10], green methods, like microwave assisted condensation [11], isomerization of *N*-tosylaziridines by palladium [12], tellurium metal assisted oxidation of chloramine T [9], rearrangement of oxime *O*-sulfonates by in situ [13, 14], reaction of *N*-trimethylsilylaldimine and sulfonyl chlorides [15] and condensation of racemic sulfinylamides with carbonyl compounds followed by oxidation with haloperoxobenzoic acid [16]. For this condensation various catalysts have been utilized such as, TiCl₄ [17], aq. HCOOH and Na₂CO₃ [18], Molecular sieves-amberlite [19], clay mineral [20], zeolites [7], Lewis acid, protic acids, base, metal oxides, AlCl₃ [21], BF₃-Et₂O [22], sulfamic acid [23],

P₂O₅/SiO₂ [24], tetraethyl orthosilicate [25], silphox [26], zirconium sulphate-silica [27], sulfated titania [28], montmorillonite K-10 [29], zirconium oxide-persulfate [30], powdered zinc oxide [31], TFFA [32], InCl₃ [10] and ionic liquids [33]. This *E-N*-benzene sulfonylimines have been used for synthesis of many organics such as triheteroarylmethanes [2], stereoselective heterocycles [1], co-ordination compounds [34], porphyrins [35], azetidines-2-imines [36], pyrrolidines [37], multisubstituted imidazolidines [38], α,β -diamino acid derivatives [39], aziridines [40], Chiral aziridines [41], (*Z*)-*N*-(1-bromo-1-alken-2-yl)-*p*-toluenesulfonamides [3], Tosylamino alkyl naphthalenols [42], vinyl aziridines [43], 2-Benzoylfumarates and 1-Aza dienes [44], complexes containing nitroamines [45], arylglycine derivatives [46], 2-alkylidenazetidines [47], chiral 1-aryl-1-ethylamines [48], heteroaromatic *N*-tosyl α -amino acids [49], carboxylates [50], tosylhexahydropyrimidine-2-imines [51], β -aryl- β -amino acids [52], furyl sulphonamides [53], Baylis-Hillman adducts [8], imidazoles [11], and sulphonamide phosphonates [54]. Solvent free methods are very useful for synthesis of various organics [11, 55-57].

Organic Chemists and researcher have paid more attention to green synthesis, due to the operational simplicity, lesser times, high yields, easy handling procedure and less hazardous to reaction and environment. The spectroscopic data is useful for analysis ground state molecular equilibration and configuration of organic molecules [58-61]. Also these findings were carried out via spectral correlations. The spectral group frequencies were correlated various substituent constants and *F* and *R* parameters gave the effects of the substituents on the group frequency of the corresponding molecule.

This type of study was studied only in ketones, esters, imines, acyl halides, pyrazolines, unsaturated aldehydes and acids [62-66]. With a view to above, there is no report available in the literature for solvent free synthesis and the study of effects of substituent on the group frequencies of some *E-N*-benzene sulfonylimines by SiO₂-H₃PO₄ catalyzed condensation of aldehydes and sulphonamides. Therefore the authors have taken efforts to synthesis of some *E-N*-benzene sulfonylimines and studied the spectral correlations on the specified group frequencies using Hammett correlation equation.

2. EXPERIMENTAL

2. 1. Materials and Methods

All chemicals used were purchased from Sigma-Aldrich and E-Merck chemical companies. Infrared spectra (KBr, 4000-400 cm⁻¹) were recorded on AVATAR-300 Fourier transform spectrophotometer. The NMR spectra of all *E-N*-benzene sulfonylimines were recorded in INSTRUM AV300 NMR spectrometer operating at 500 MHz for ¹H and 125.46 MHz for ¹³C spectra in DMSO solvent using TMS as internal standard. Mass spectra were recorded on a SIMADZU GC-MS2010 Spectrometer using Electron Impact (EI) techniques.

2. 2. Synthesis of SiO₂-H₃PO₄ catalyst

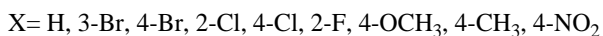
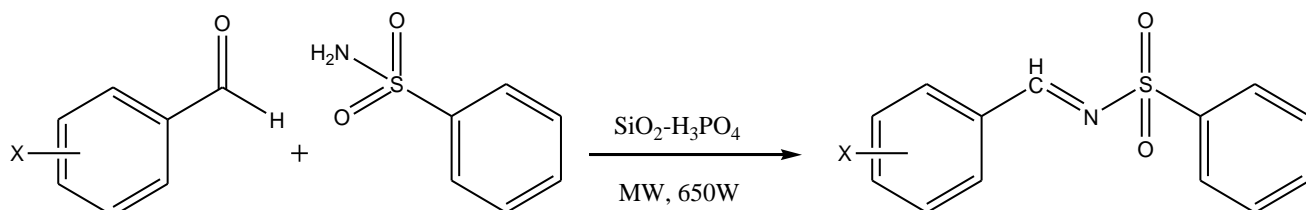
The SiO₂-H₃PO₄ catalyst was prepared by procedure published in literature [55].

2. 3. Synthesis of *E-N*-benzene sulfonylimines

To an appropriate mixture of substituted benzaldehydes (2 mmol), benzene sulphonamide (2 mmol) and 0.5 g of SiO₂-H₃PO₄ were taken in 50 mL glass beaker and covered with lid. This mixture was subjected to microwave heating for 5-8 minutes at 650 W

(Samsung GW73BD microwave oven, 100-750 W, 230 V A/C, 2450 MHz). The completion of the reaction was monitored by thin layer chromatography. After completion of reaction, dichloromethane (20 mL) was added, followed by simple filtration.

The pure product was obtained by the evaporation of dichloromethane and recrystallization with ethanol. The yields of the *E*-*N*-benzene sulfonilimines are more than 90 %. The purities of these *E*-*N*-benzene sulfonilimines were persuaded by their physical constants, IR, ^1H , ^{13}C NMR and Mass spectral data. Analytical and mass spectral data are presented in Table 1.



Scheme 1. Synthesis of substituted aryl *E*-*N*-benzene sulfonilimines.

Table 1. Analytical, Physical constants and mass fragment data of *E*-*N*-benzene sulfonilimines.

Entry	X	M. F.	M. W.	Yield (%)	m. p. (°C)	Mass (m/z)
1	H	C ₁₃ H ₁₁ NO ₂ S	245	94	83-84 (80-82)[33]	245[M ⁺]
2	3-Br	C ₁₃ H ₁₀ BrNO ₂ S	325	91	111-112	324[M ⁺], 326[M ⁺²]
3	4-Br	C ₁₃ H ₁₀ BrNO ₂ S	325	93	199-200 (198-204)[33]	325[M ⁺], 327[M ⁺²]
4	2-Cl	C ₁₃ H ₁₀ ClNO ₂ S	279	91	135-136 (132-136)[33]	279[M ⁺], 281[M ⁺²]
5	4-Cl	C ₁₃ H ₁₀ ClNO ₂ S	279	91	129-130 (128-132)[33]	279[M ⁺], 281[M ⁺²]
6	2-F	C ₁₃ H ₁₀ FNO ₂ S	263	90	124-125 (119-125)[33]	263[M ⁺], 265[M ⁺²]
7	4-OCH ₃	C ₁₄ H ₁₃ NO ₃ S	275	94	133-134 (130-134)[33]	275[M ⁺]
8	4-CH ₃	C ₁₄ H ₁₃ NO ₂ S	259	92	112-113 (110-112)[33]	259[M ⁺]
9	4-NO ₂	C ₁₄ H ₁₃ N ₂ O ₄ S	290	90	113-114 (110-112)[33]	290[M ⁺]

3. RESULTS AND DISCUSSION

3. 1. Spectral linearity

In the present study the Hammett spectral linearity of these synthesised *E-N*-benzene sulfonylimines has been studied by evaluating the substituent effects [61-66] on the group frequencies. The assigned spectroscopic data of all *E-N*-benzene sulfonylimines such as absorption infrared carbonyl stretches of $\nu\text{C}=\text{N}$, $\nu\text{S}=\text{O}$, (cm^{-1}), NMR chemical shifts $\delta(\text{ppm})$ of CH, C=N have been assigned and these data are correlated with various substituent constants.

3. 1. 1. IR spectral study

The assigned C=N and S=O frequencies (cm^{-1}) of synthesised *E-N*-benzene sulfonylimines of present study are tabulated in Table 2. These data have been correlated with Hammett substituent constants [55-66] and Swain-Lupton constants [67] and are presented in Table 3. In this correlation the structure parameter Hammett equation employed is as shown in the following equation:

$$\nu = \rho\sigma + \nu_0 \quad \dots (1)$$

Where ν is the carbonyl frequencies of substituted system and ν_0 is the corresponding quantity of unsubstituted system; σ is a Hammett substituent constant, which in principle is characteristics of the substituent and ρ is a reaction constant which is depend upon the nature of the reaction.

Table 2. Infrared stretches (ν , cm^{-1}) and NMR chemical shifts (δ , ppm) of *E-N*-benzene sulfonylimines.

Entry	X	IR		¹ H	¹³ C
		$\nu\text{C}=\text{N}$	$\nu\text{S}=\text{O}$	$\delta\text{C-H}$	$\delta\text{C}=\text{N}$
1	H	1658.25	1063.58	8.462	179.32
2	3-Br	1657.98	1060.35	8.495	178.26
3	4-Br	1658.04	1062.34	8.359	179.55
4	2-Cl	1658.94	1062.52	8.401	179.02
5	4-Cl	1657.92	1062.85	8.442	179.46
6	2-F	1658.03	1062.38	8.325	179.65
7	4-OCH ₃	1657.02	1060.02	8.001	178.02
8	4-CH ₃	1658.07	1060.21	8.012	178.98
9	4-NO ₂	1659.03	1064.25	8.499	181.36

The results of single parameter statistical analysis of $\nu\text{C}=\text{N}$ stretches with Hammett σ and σ^+ substituent constants gave satisfactory correlations. The Hammett σ_I , σ_R and R parameters were satisfactorily correlated with $\nu\text{C}=\text{N}$ stretches excluding H, 2-Cl, 2-F and 4-CH₃ substituents. If these substituents were included in the correlation, the correlations were reduced significantly. The field effect components of the substituents were fail in correlation. All correlations gave positive ρ values. This may mean that the normal substituent effect operates in all *E-N*-benzene sulfonylimines. The failure in correlation is due the conjugation between the substituent and the C=N group in all *E-N*-benzene sulfonylimines as shown in Figure 1.

The correlation of $\nu_{\text{C=S}}$ stretches with Hammett σ , σ^+ and σ_{I} substituent constants gave satisfactory r values [σ ($r = 0.906$), σ^+ ($r = 0.906$) and σ_{I} ($r = 0.903$)] excluding H, 3-Br and 4-CH₃ substituents. If these substituents were included in the correlation, the correlations were reduced significantly. The resonance and field effect components of the substituents were fail in correlations. While seeking the ρ values, all correlation gave positive ρ values and it is evident for the normal substituent effects operates in all systems. The failure in correlation is due the conjugation between the substituent and the C=N group in all *E-N*-benzene sulfonylimines as shown in Figure 1.

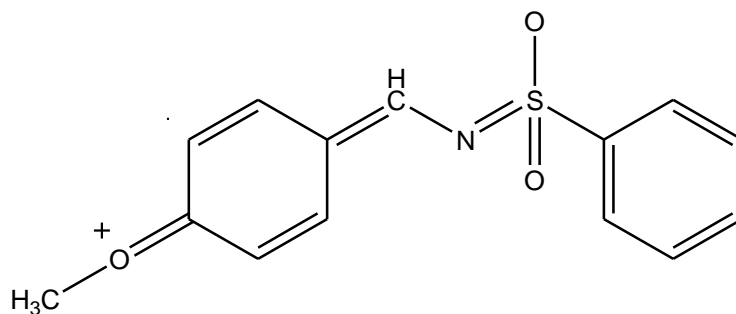


Figure 1. The resonance –conjugative structure.

In view of the inability of some of the σ constants to produce individually satisfactory correlations, it was thought that worthwhile to seek multiple correlations involving either σ_{I} and σ_{R} constants or Swain-Lupton's [67] F and R parameters. The correlation equations for $\nu_{\text{C=N}}$ and $\nu_{\text{C=S}}$ are given in equations 2-5.

$$\nu_{\text{C=N}}(\text{cm}^{-1}) = 1658.14(\pm 0.342) + 0.902(\pm 0.071) \sigma_{\text{I}} + 1.875(\pm 0.081) \sigma_{\text{R}} \quad \dots(2)$$

$(R = 0.969, n = 9, P > 95 \%)$

$$\nu_{\text{C=N}}(\text{cm}^{-1}) = 1658.21(\pm 0.281) + 1.000(\pm 0.060) F + 1.965(\pm 0.654) R \quad \dots(3)$$

$(R = 0.979, n = 9, P > 95 \%)$

$$\nu_{\text{C=S}}(\text{cm}^{-1}) = 1061.88(\pm 0.957) + 2.468(\pm 0.243) \sigma_{\text{I}} + 4.097(\pm 1.249) \sigma_{\text{R}} \quad \dots(4)$$

$(R = 0.963, n = 9, P > 95 \%)$

$$\nu_{\text{C=S}}(\text{cm}^{-1}) = 1061.99(\pm 0.811) + 2.839(\pm 1.446) F + 4.455(\pm 1.888) R \quad \dots(5)$$

$(R = 0.972, n = 9, P > 95 \%)$

3. 1. 2. ¹H NMR spectral study

Deuterated dimethylsulphoxide was used for recording the ¹H NMR spectra of synthesized *E-N*-benzene sulfonylimines employing tetramethylsilane (TMS) as internal standard. The CH proton signals of the *E-N*-benzene sulfonylimines were obtained as a singlet and assigned from their spectra. The assigned CH proton chemical shifts δ (ppm) of all *E-N*-benzene sulfonylimines were presented in Table 2.

Table 3. Results of statistical analysis of infrared absorptions (ν , cm^{-1}) and NMR chemical shifts (δ , ppm) of *E-N*-benzene sulfonilimines with Hammett σ , σ^+ , σ_I , σ_R and F and R parameters.

Frequency	Constants	r	I	ρ	s	n	Correlated derivatives
$\nu\text{C}=\text{N}$ (cm^{-1})	σ	0.906	1657.90	1.295	0.45	9	H, 3-Br, 4-Br, 2-Cl, 4-Cl, 2-F, 4-OCH ₃ , 4-CH ₃ , 4-NO ₂
	σ^+	0.907	1657.97	0.941	0.45	9	H, 3-Br, 4-Br, 2-Cl, 4-Cl, 2-F, 4-OCH ₃ , 4-CH ₃ , 4-NO ₂
	σ_I	0.903	1657.84	0.824	0.59	7	3-Br, 4-Br, 2-Cl, 4-Cl, 2-F, 4-OCH ₃ , 4-NO ₂
	σ_R	0.905	1658.46	1.819	0.50	7	H, 3-Br, 4-Br, 4-Cl, 4-OCH ₃ , 4-CH ₃ , 4-NO ₂
	F	0.802	1657.92	0.575	0.96	9	H, 3-Br, 4-Br, 2-Cl, 4-Cl, 2-F, 4-OCH ₃ , 4-CH ₃ , 4-NO ₂
	R	0.906	1658.53	1.713	0.46	7	H, 3-Br, 4-Br, 4-Cl, 4-OCH ₃ , 4-CH ₃ , 4-NO ₂
$\nu\text{S}=\text{O}$ (cm^{-1})	σ	0.906	1061.51	2.984	1.28	7	4-Br, 2-Cl, 4-Cl, 2-F, 4-OCH ₃ , 4-CH ₃ , 4-NO ₂
	σ^+	0.906	1061.69	2.021	1.25	7	4-Br, 2-Cl, 4-Cl, 2-F, 4-OCH ₃ , 4-CH ₃ , 4-NO ₂
	σ_I	0.903	1061.22	2.297	1.52	7	H, 3-Br, 4-Br, 2-Cl, 4-Cl, 2-F, 4-OCH ₃ , 4-NO ₂
	σ_R	0.851	1062.75	3.946	1.41	9	H, 3-Br, 4-Br, 2-Cl, 4-Cl, 2-F, 4-OCH ₃ , 4-CH ₃ , 4-NO ₂
	F	0.831	1061.34	1.875	1.55	9	H, 3-Br, 4-Br, 2-Cl, 4-Cl, 2-F, 4-OCH ₃ , 4-CH ₃ , 4-NO ₂
	R	0.805	1062.91	3.739	1.34	9	H, 3-Br, 4-Br, 2-Cl, 4-Cl, 2-F, 4-OCH ₃ , 4-CH ₃ , 4-NO ₂
δCH (ppm)	σ	0.907	8.245	0.485	0.12	8	H, 3-Br, 4-Br, 2-Cl, 4-Cl, 2-F, 4-OCH ₃ , 4-CH ₃
	σ^+	0.908	8.275	0.322	0.12	8	H, 3-Br, 4-Br, 2-Cl, 4-Cl, 2-F, 4-OCH ₃ , 4-CH ₃
	σ_I	0.905	8.173	0.444	0.17	7	3-Br, 4-Br, 2-Cl, 4-Cl, 2-F, 4-CH ₃ , 4-NO ₂
	σ_R	0.845	8.414	0.461	0.18	9	H, 3-Br, 4-Br, 2-Cl, 4-Cl, 2-F, 4-OCH ₃ , 4-CH ₃ , 4-NO ₂
	F	0.804	8.210	0.320	0.18	9	H, 3-Br, 4-Br, 2-Cl, 4-Cl, 2-F, 4-OCH ₃ , 4-CH ₃ , 4-NO ₂
	R	0.905	8.435	0.444	0.17	7	H, 3-Br, 4-Br, 2-Cl, 4-Cl, 2-F, 4-NO ₂
$\delta\text{C}=\text{N}$ (ppm)	σ	0.907	178.88	2.235	0.69	8	H, 4-Br, 2-Cl, 4-Cl, 2-F, 4-OCH ₃ , 4-CH ₃ , 4-NO ₂
	σ^+	0.906	179.05	1.333	0.75	8	H, 4-Br, 2-Cl, 4-Cl, 2-F, 4-OCH ₃ , 4-CH ₃ , 4-NO ₂
	σ_I	0.904	178.66	1.744	0.92	7	3-Br, 4-Br, 2-Cl, 4-Cl, 2-F, 4-OCH ₃ , 4-NO ₂
	σ_R	0.905	179.79	2.860	0.83	8	H, 3-Br, 4-Br, 2-Cl, 4-Cl, 4-OCH ₃ , 4-CH ₃ , 4-NO ₂
	F	0.844	178.65	1.679	0.92	9	H, 3-Br, 4-Br, 2-Cl, 4-Cl, 2-F, 4-OCH ₃ , 4-CH ₃ , 4-NO ₂
	R	0.855	179.83	2.362	0.84	9	H, 3-Br, 4-Br, 2-Cl, 4-Cl, 2-F, 4-OCH ₃ , 4-CH ₃ , 4-NO ₂

r = correlation coefficient; I = intercept; ρ = slope; s = standard deviation;
n = number of correlated derivative

In nuclear magnetic resonance spectra, the proton or the ^{13}C chemical shifts (δ) depends on the electronic environment of the nuclei concerned. The assigned vinyl proton chemical shifts (ppm) have been correlated with reactivity parameters using Hammett equation in the form of

$$\text{Log } \delta = \text{Log } \delta_0 + \rho\sigma \quad \dots (6)$$

where δ_0 is the chemical shift of unsubstituted ketones.

The assigned CH proton chemical shifts (ppm) are correlated with Hammett σ constants and F and R parameters [55-66]. The results of statistical analysis [55-66] are presented in Table 3. The obtained correlations were satisfactory for CH chemical shifts with Hammett σ , σ^+ , σ_I substituent constants and R parameters excluding H, 4-OCH₃, 4-CH₃ and 4-NO₂ substituents. The Hammett σ_R constants and R parameters gave poor correlation. All correlation gave positive ρ values and it is evident for the normal substituent effects operates in all correlations. The failure in correlation was due to the weak and incapable of resonance and field effects of substituents for prediction of the effects on the CH chemical shifts in all systems and it is associated with the resonance-conjugates structure shown in Figure 1.

Application of Swain-Lupton's [67] treatment to the relative chemical shifts (δ , ppm) of CH with F and R values is successful with resonance, inductive effect generates the multi regression equations 7-8.

$$\delta_{\text{CH}}^{(\text{ppm})} = 8.252(\pm 0.107) - 0.462(\pm 0.220) \sigma_I + 0.489(\pm 0.281) \sigma_R \quad \dots (7)$$

($R = 0.972$, $n = 9$, $P > 95\%$)

$$\delta_{\text{CH}}^{(\text{ppm})} = 8.0292(\pm 0.096) - 0.440(\pm 0.027) F + 0.555(\pm 0.222) R \quad \dots (8)$$

($R = 0.976$, $n = 9$, $P > 95\%$)

3. 1. 3. ^{13}C NMR spectral study

Physical organic chemists, researchers, organic chemists, spectral analysts, and scientists [55-66] have made extensive study of ^{13}C NMR spectral correlations of ketones, enones, acyl chlorides, bromides, esters, imines, pyrazolines and styrenes. The assigned vinyl C=N carbon chemical shifts are presented in Table 2. The results of statistical analysis are given in Table 3.

The C=N chemical shifts (ppm) gave satisfactory correlation with Hammett substituent constants excluding 3-Br, 2-F and 4-CH₃ substituents. All correlations produced positive ρ values. This means that the normal substituent effect operates in all correlations. The F and R parameters were fail in correlation. This is due to reasons stated earlier with the resonance conjugative structure shown in Figure 1.

The Swain Lupton's [67] parameter correlations were satisfactorily obtained within these carbon chemical shifts and the regression equations are given in 9-10.

$$\delta_{\text{C=N}}^{(\text{ppm})} = 179.14(\pm 0.521) + 1.867(\pm 1.111) \sigma_I + 2.976(\pm 1.362) \sigma_R \quad \dots (9)$$

($R = 0.973$, $n = 9$, $P > 95\%$)

$$\delta_{\text{C=N}}^{(\text{ppm})} = 179.08(\pm 0.423) - 2.317(\pm 0.911) F + 2.946(\pm 0.948) R \quad \dots (10)$$

($R = 0.982$, $n = 9$, $P > 95\%$)

4. CONCLUSIONS

We have synthesised more than 90 % yield of some *E-N*-benzene sulfonylimines using a solid SiO₂-H₃PO₄ acid catalyst by solvent free condensation of benzene sulfonyl chloride and substituted benzaldehydes under microwave irradiation. The effects of substituent on the group frequencies (ν , cm⁻¹) such as C=N, S=O and the chemical shifts (δ , ppm) of CH and C=N all the *E-N*-benzene sulfonylimines have been studied.

REFERENCES

- [1] Weinreb S. M., *Topics in Curr. Chem.* 190 (1997) 131-184.
- [2] Temelli B., Tasgin D. I., Unaleroglu C., *Tetrahedron* 66(34) (2010) 6765-6768.
- [3] Yamagishi M., Nishigai K., Hata T., Urabe H., *Org. Lett.* 13(18) (2011) 4873-4875.
- [4] Ma Z., Zhao Y., Jiang N., Jin X., Wang J., *Tetrahedron Lett.* 43 (2002) 3209-3212.
- [5] Hojo M., Murakami C., Fujii A., Hosomi A., *Tetrahedron Lett.* 40 (1999) 911-914
- [6] Yamada K. I., Yamamoto Y., Maekawa M., Tomioka K., *J. Org. Chem.* 69 (2004) 1531-1534.
- [7] Wang K., Xing Z., Ma Y., *Catal. Lett.* 123 (2008) 129-134.
- [8] Shi Y. L., Xu Y. M., Shi M., *Advanced Synth. Catal.* 346 (9-10) (2004) 1220-1230.
- [9] Trost B. M., Marrs C., *J. Org. Chem.* 56 (1991) 6468-6470.
- [10] Deng G. S., Zou J. Y., Sun T. F., *Chin. Chem. Lett.* 22(5) (2011) 511-514.
- [11] Samanta S. K., Kylänlahti I., Yli-Kauhaluoma J., *Bioorg Med Chem Lett.* 15(16) (2005) 3717-3719.
- [12] Wolfe J. P., Ney J. E., *Org. Lett.* 5 (2003) 4607-4710.
- [13] Boger D. L., Corbett W. L., *J. Org. Chem.* 57 (1992) 4777-4780 .
- [14] Sisko J., Weinreb S. M., *J. Org. Chem.* 55 (1990) 393-395.
- [15] Georg G. I., Harriman G. C. B., Peterson S. A., *J. Org. Chem.* 60 (1995) 7366-7368.
- [16] Ruano J. L. G., Aleman J., Cid M. B., Parra A., *Org. Lett.* 7 (2005) 179-182 .
- [17] Wynne J. H., Price S. E., Rorer J. R., *Synth. Commun.* 33 (2003) 341-342.
- [18] Chemla F., Hebbe V., Normant J. F., *Synthesis.* (2000) 75-76.
- [19] Vishwakarma L. C., Stringer O. D., Davis F., *Org Synth Coll VIII* (1993) 546-549.
- [20] Laszlo P., *Science* 235 (1987) 1473-1477.
- [21] Sharghi H., Hosseini-Sarvari M., Ebrahimpourmoghaddam S., *Arkivoc*, 15 (2007) 255-264.
- [22] McFarlane A. K., Thomas G., Whiting A., *Tetrahedron Lett.* 34 (1993) 2379-2382.
- [23] Li Z., Ren X., Wei P., Honggui Wan, Shi Y., Ouyang P., *Green Chem.* 8 (2006) 433-436.
- [24] Hasaninejad A., Zare A., Shrgi H., *Arkivoc.* 11 (2008) 64-74.

- [25] Love B. E., Raje P. S., Willims T. C., *Synlett.* (1994) 493-494.
- [26] Hasaninejad A., Zare J., *Sulfur Chem.* 28 (2007) 357-364.
- [27] Geng L. J., Liu Y., *Chin. J. Chem.* 25 (2005) 1103-1106.
- [28] Jin T. S., Feng G. L., Yang M. N., *J. Chem. Res (S)*. 9 (2003) 591-594.
- [29] Jin T. S., Yu M. J., Liu L. B., *Synth. Commun.* 36 (2006) 2339-2344.
- [30] Jin T. S., Feng G. L., Yang M. N., *Synth. Commun.* 34 (2004) 1277-1283.
- [31] Hoseini-Sarvari M., Sharghi H., *Phosphorus, Sulfur and Silicon.* 182 (2007) 2125-2130.
- [32] Lee K. Y., Lee C. G., Kim J. N., *Tetrahedron Lett.* 44 (2003) 1231-1234.
- [33] Ardhapure S. S., Siddiqui S. A., Tekales U., Ingle R. D., Shinde S. B., Pawar R. P., *Eur. Chem. Bull.* 2(6) (2013) 320-323.
- [34] Gao F., Deng M., Qian C., *Tetrahedron.* 61(52) (2005) 12238-12243.
- [35] Temelli B., Unaleroglu C., *Tetrahedron.* 65(10) (2009) 2043-2050.
- [36] Van Camp A., Goossens D., Moya-Portugues M., Marchand-Brynaert J., Ghosez L., *Tetrahedron Lett.* 21(32) (1980) 3081-3084.
- [37] Meng L. G., CAi P., Guo Q., Xue S., *J. Org. Chem.* 73(21) (2008) 84-91-8496.
- [38] Wu X., Zhang J., *Synthesis.* 44(14) (2012) 2147-2154.
- [39] Song G., Jin M., Li Z., Ouyanf P., *Org. Biomol. Chem.* 9 (2011) 7144-7150.
- [40] Concellon J. M., Rodriguez-Solla H., Simal C., *Org. Lett.* 10 (2008) 4457-4460.
- [41] Hayashi Y., Urushima T., Sakamoto D., Torri K., Ishikawa H., *Chem.- A European J.* 17(42) (2011) 11715-11718.
- [42] Ravindra Reddy .C, Sindhu C., Nandan Kumar D., Krishnaiah M., *Helvetica Chim, Acta.* 94 (2) (2011) 289-292.
- [43] Hou X-L., Yang X-F., Dai L-X., *Chem. Commun.* (1998) 747-748.
- [44] Nair V., Sreekanth A. R., Abhilash N., Biju A. T., Reena Devi B., Menon R. S., Rath N. P., Srinivas R., *Synthesis.* (2003) 1895-1902.
- [45] Zhou H., Peng D., Qin B., Hou Z., Liu X., Feng X., *J. Org. Chem.* 72(26) (2007) 10302-10304.
- [46] Yamamoto Y., Takashi Y., Kurihara K., Miyaura N., *Australian J. Chem.* 64(11) (2011) 1447-1453.
- [47] Lu H., Li C., *Org. Lett.* 8 92006) 5365-5367.
- [48] Nishimura T., Yasuhara Y., Hayashi T., *Org. Lett.* 8 (2006) 979-981.
- [49] Johannsen M., *Chem. Commun.* (1999) 2233-2234.
- [50] Lu K., Kwon O., *Org. Synthesis.* 86 (2009) 212-220.
- [51] Shutalev A. D., Fesenko A. A., *Tetrahedron.* 67(36) (2011) 6883-6888.
- [52] Ma Z., Zhao Y., Jiang N., Jin X., Wang J., *Tetrahedron Lett.* 43(17) (2002) 3209-3214.

- [53] Pawda A., Zanka A., Cassidy M. P., *Tetrahedron*. 59(27) (2003) 4939-4944.
- [54] Das B., Balasubramaniam P., Krishnaiah M., Veeranjanyulu B., Reddy G. C., *J. Org. Chem.* 74(11) (2009) 4393-4395.
- [55] Janaki P., Sekar K. G., Thirunarayanan G., *Org. Chem: An Indian J.* 9(2) (2013) 68-80.
- [56] Thirunarayanan G., Vanangamudi G., Subramanian M., *Org. Chem: An Indian J.* 9(1) (2013) 1-16.
- [57] Sakthinathan S. P., Suresh R., Mala V., Sathiyamoorthi K., Kamalakkannan D., Ranganathan K., Arulkumaran R., Vijayakumar S., Sundararajan R., Vanangamudi G., Thirunarayanan G., *Int. Lett. Chem. Phys. Astro.* 6 (2013) 77-90.
- [58] Sathiyamoorthi K., Mala V., Suresh R., Sakthinathan S. P., Kamalakkannan D., Ranganathan K., Arulkumaran R., Sundararajan R., Vijayakumar S., Vanangamudi G., Thirunarayanan G., *Int. Lett. Chem. Phys. Astro.* 7(2) (2013) 102-119.
- [59] Sekar K. G., Thirunarayanan G., *Int. Lett. Chem. Phys. Astro.* 8(2) (2013) 160-174.
- [60] Arulkumaran R., Vijayakumar S., Sundararajan R., Sakthinathan S. P., Kamalakkannan D., Suresh R., Ranganathan K., Vanangamudi G., and Thirunarayanan G., *Int. Lett. Chem. Phys. Astro.* 4 (2012) 17-38.
- [61] Ranganathan K., Suresh R., Kamalakkannan D., Arulkumaran R., Sundararajan R., Sakthinathan S. P., Vijayakumar S., Vanangamudi G., Thirumurthy K., Mayavel P., Thirunarayanan G., *Int. Lett. Chem. Phys. Astro.* 4 (2012) 66-75.
- [62] Subramanian M., Vanangamudi G., Thirunarayanan G., *Spectrochim Acta* 110A (2013) 116-123.
- [63] Arulkumaran R., Vijayakumar S., Sundararajan R., Sakthinathan S. P., Kamalakkannan D., Suresh R., Ranganathan K., Rajakumar P. R., Vanangamudi G., Thirunarayanan G., *Int. Lett. Chem. Phys. Astro.* 5 (2013) 21-38.
- [64] Thirunarayanan G., Vanangamudi G., Sathiyendiran V., Ravi K., *Indian J. Chem.* 50B (4) (2011) 593-604.
- [65] Suresh .R, Kamalakkannan D., Ranganathan K., Arulkumaran R., Sundararajan R., Sakthinathan S. P., Vijayakumar S., Sathiyamoorthy K., Mala V., Vanangamudi G., Thirumurthy K., Mayavel P., Thirunarayanan G., *Spectrochim. Acta* 101A (2013) 239-248.
- [66] Sakthinathan S. P., Vanangamudi G., Thirunarayanan G., *Spectrochim. Acta* 95A (2012) 693-700.
- [67] Swain C. G., Lupton E. C. Jr., *J. Am. Chem. Soc.* 90 (1968) 4328-4337.