

Photocatalytic mineralisation of humic acids using TiO₂ modified by tungsten dioxide/ hydrogen peroxide

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TiO₂ of anatase structure was modified by tungsten dioxide and H₂O₂ in order to obtain WO₃-TiO₂ photocatalyst with enhanced photocatalytic activity under both, UV and artificial solar light irradiations. WO₃ was dissolved in 30% H₂O₂ and mixed with TiO₂ in a vacuum evaporator at 70°C. Such modified TiO₂ was dried and then calcinated at 400 and 600°C.

The prepared samples and unmodified TiO₂ were used for the photocatalytic decomposition of humic acids (Leonardite standard IHSS) in the aqueous solution under irradiations of both, UV and artificial solar light. Modification of TiO₂ with tungsten dioxide and H₂O₂ improved separation of free carriers in TiO₂ which resulted in the increase of OH radicals formation. Calcination caused an increase of anatase crystals and higher yield in OH radicals. The uncalcined samples showed high abilities for the adsorption of HA. Combination of adsorption abilities and photocatalytic activity of photocatalyst caused that the uncalcined TiO₂ modified with WO₃/H₂O₂ showed the shortest time of HA mineralisation. The mineralisation of HA under the artificial solar light was much lower than under the UV. It was proved that, although OH radicals are powerful in the decomposition of HA, adsorption can facilitate the contact of the adsorbed molecules with the photocatalyst surface and accelerate their photocatalytic decomposition.

Keywords: WO₃/TiO₂, humic acids photodecomposition, OH radicals formation.

INTRODUCTION

Humic acids naturally present in the surface water state a problem because of their abilities to form the trihalomethanes (THMs) during the chlorination of water. THMs are classified as carcinogenic compounds¹. The removal of humic acids by the conventional methods such as coagulation or filtration is not completely satisfactory, therefore the other solutions for HA destruction are searched and applied². It has already been reported that humic acids (HA) could be efficiently removed from water through the photocatalytic process, which belongs to Advances Oxidation Processes³⁻⁸. Formed in these processes the hydroxyl radicals are strong oxidising species, which are responsible for the decomposition of organic matter in water. The OH radicals can be generated by ozonation, photolysis of H₂O₂, photo-Fenton reaction or photocatalysis, among these processes the last one is interesting because of the low cost of the photocatalyst used and the relative high efficiency. The most common photocatalyst tested for the photocatalytic decomposition of HA is TiO₂⁴⁻⁸. In a common photocatalytic process the UV irradiation is used for the excitation of TiO₂, however solar light with a small content of UV rays (around 4%) is also applied. It was reported that artificial sunlight irradiation was not sufficient to cause a significant decomposition of HA, whereas TiO₂ appeared to be very effective under the UV irradiation⁸. The macromolecular humic acids on the visible light-illuminated TiO₂ are degraded with the reduction of aromatic character but not to the complete mineralization⁶. Many researchers noticed that HA are strongly adsorbed on the surface of TiO₂, which has influence on the surface oxidative mechanism⁴⁻⁸. Some authors reported that the photodegradation of HA could proceed on the surface via the oxidation by hole and this mechanism was favoured in the case of the

adsorbed molecules⁵. HA was progressively degraded by the surface oxidative mechanism and so long as some macromolecules issued from HA remained in solution this mechanism predominated⁵. Some researchers prepared the composite of TiO₂ and granular activated carbon (GAC) in order to enhance the adsorption abilities of the photocatalyst. They reported that the degradation of HA on the TiO₂/GAC composite was facilitated by the synergistic relationship between surface adsorption characteristics and photocatalytic potential⁷.

The disadvantage of using the TiO₂ photocatalyst is that only UV light can be applied for its excitation, because it has a relatively high value of the band gap energy, around 3.2 eV (for active anatase phase). The efficiency of TiO₂ photocatalyst can be increased by doping some metals as silver, gold or platinum, which are electron acceptors and can retard the inconvenient recombination reaction, which occurs in TiO₂ after its excitation, giving better separation of the free carriers and increasing the OH radicals formation⁹⁻¹¹. Another solution of better photocatalytic performance is the preparation of a new generation of photocatalysts, which demonstrates activity under visible light. A lot of work was done to prepare non metal doping TiO₂ (incorporation of nitrogen, carbon, sulphur atoms, etc.)¹²⁻¹⁶.

The other possibility is the preparation of a mixture of two semiconductors with different energies of the band gap, such as TiO₂ doping with WO₃. WO₃ has the value of the band gap energy about 2.8 eV and can absorb the light in the visible range. It was reported by Li and Song¹⁷⁻¹⁸ that when the coupled photocatalyst WO₃/TiO₂ was prepared by the sol-gel method, no stoichiometric solid solution of W_xTi_{1-x}O₂ would be formed inside the forbidden band of TiO₂ which induces a tungsten impurity energy level responsible for the activity under visible light. Shifu et al.¹⁹ reported

that when WO_3 and TiO_2 form a coupled photocatalyst, TiO_2 and WO_3 can be excited simultaneously under UV illumination. As the conduction band of WO_3 is lower than that of TiO_2 , the former can act as a sink for the photogenerated electrons. The photogenerated electrons of the TiO_2 conduction band will be transferred to the conduction band of WO_3 . Since the holes move in the opposite direction from the electrons, photogenerated holes might be trapped within the TiO_2 particle, which makes charge separation more efficient^{19–25}. A lot of methods of WO_3/TiO_2 preparation have been used and reported on, such as impregnation, sol-gel, hydrothermal, sol-precipitation, ball milling, electro-deposition, metal ion implantation, etc.^{18–30}.

In the presented manuscript photodecomposition of HA was performed on the prepared WO_3/TiO_2 photocatalyst. For the preparation of WO_3/TiO_2 the impregnation method was selected in order to obtain a good dispersion of tungsten oxide particles on the surface of TiO_2 . Well dispersed WO_3 particles on the surface of TiO_2 can sufficiently retard the recombination reaction and increase the absorption of the visible light. We used this method, because we would like to use for the preparation a raw industrially produced TiO_2 material, which has very high BET surface area, around $300 \text{ m}^2/\text{g}$. The disadvantage of this method consists in the fact that it does not allow the narrowing of the band gap by introducing the tungsten oxide phase inside the forbidden band of TiO_2 , as it is possible using the sol-gel, therefore excitation under visible light is only possible in the WO_3 semiconductor. However, under UV-Vis irradiation both, the WO_3 and TiO_2 semiconductors can be excited and then the transfer of the electrons from the conductive band of TiO_2 to WO_3 is possible, making the separation of free carriers more efficient, which increases the yield of photochemical reactions.

The photoactivity of the prepared WO_3/TiO_2 photocatalysts under both, UV and artificial solar light irradiations towards HA decomposition is discussed. The theoretical amount of WO_3 in TiO_2 was 3%. This was an optimal value selected on the basis of the previous experiments, also being in a good agreement with the literature data^{17,19}.

EXPERIMENTAL

Materials

TiO_2 was supplied by the Chemical Factory Police S.A. in Poland. It consists of the anatase phase, around 11.5 wt % and a small amount of rutile phase, about 3.5 wt %, and has high BET surface area, around $275 \text{ m}^2/\text{g}$. WO_2 purchased from Sigma-Aldrich Co. (a.g.) was used as a tungsten precursor. 30% H_2O_2 purchased from Scharlau Chemie was used as a dissolving agent in the preparation of WO_3/TiO_2 . Leonardite humic acid was purchased from the IHSS (International Humic Substances Society). According to the elemental analyses obtained from the list of IHSS products Leonardite HA standard (IS104H) contains in % (w/w): 63.81 C, 3.70 H, 31.27 O, 1.23 N, 0.76 S, < 0.01 P, and 2.58 ash.

Photocatalysts preparation

The WO_3/TiO_2 photocatalyst was prepared by the impregnation method. After impregnation the samples were calcinated at 400 and 600°C. The theoretical amount of the formed WO_3 in WO_3/TiO_2 was 3 wt.%. For the preparation 3 g of TiO_2 and 0.084 g of WO_2 were used. The detailed procedure was already described in the previous paper²².

Characterization of photocatalysts

The phase composition of the photocatalysts was measured by the XRD powder diffraction. The measurements were performed in the X'Pert PRO diffractometer of Philips Company, with the $\text{CuK}\alpha$ lamp (35 kW, 30mA). The obtained XRD patterns were compared with the JCPDS (Join Committee on Powder Diffraction Standards) cards.

The UV-Vis diffuse reflectance spectra were taken in the Jasco V-530 spectrometer. These spectra were recorded in the range of 250 to 800 nm with BaSO_4 as a reference and transformed to Kubelka-Munk equation for the indirect semiconductor.

The microstructure of the photocatalysts surface was analysed via a high resolution transmission electron microscope (HR-TEM)-FEI TecnaiF30.

The electrokinetic potential of the photocatalysts surface was measured by the connection of electrophoresis and LDV (Laser Doppler Velocimetry) method in Zetasizer Nano ZS.

The BET surface area was determined by the measurements of nitrogen adsorption and desorption isotherms at 77 K using Autosorb 3 (Quantachrome, U.S.A.) equipment. Prior to the adsorption measurements the samples were degassed for 24 h at 105°C under high vacuum.

The formation of OH radicals on the photocatalysts surface under both, UV and artificial solar light irradiations was measured by the fluorescence method. In this method the photocatalyst is placed into the coumarin solution, where under the irradiation of the light formed on the photocatalyst surface the OH radicals easily react with coumarin to produce a highly fluorescent product, 7-hydroxycoumarin, which was analysed in the fluorescence spectrometer Hitachi F-2500 [22, 31].

Photocatalytic activity test

The photoactivities of the prepared samples were tested for humic acids (HA) decomposition under the artificial visible light and UV irradiations.

The fluorescence lamp was used as a source of artificial visible light. The measured intensity of the incident light during photoreaction was $42 \text{ W}/\text{m}^2$ in the range of visible light (350–750 nm) and $0.8 \text{ W}/\text{m}^2$ in the range of UV. In the case of UV irradiation the Philips solar UV lamp was used. This lamp emits the light at the UV range with the intensity of $165 \text{ W}/\text{m}^2$ and at the visible region about $165 \text{ W}/\text{m}^2$. The photoemission spectra of both lamps are shown in Fig. 1.

The model solution of HA was prepared by dissolving of Leonardite HA in an ultra pure distilled water with addition of a little amount of 1 M NaOH to increase their solubility and followed filtration through the membrane filter of $0.45 \mu\text{m}$. The concentration of HA of the prepared solution was estimated on the basis of

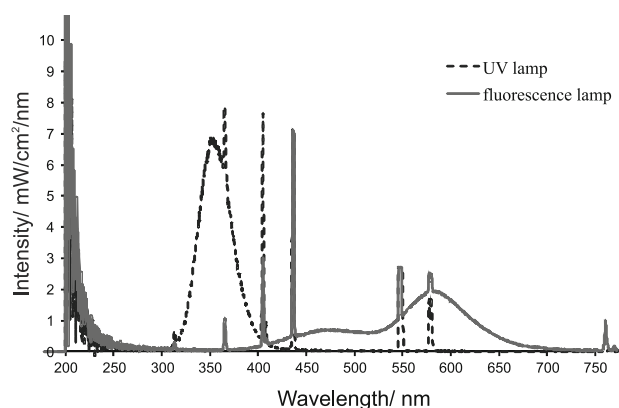


Figure 1. The photoemission spectra of UV and fluorescence lamps

TOC analyses and the calculation regarding the content of carbon in HA given by the producer.

HA degradation was performed in a batch type reactor with the suspension of 0.2 g/L photocatalysts and 0.5 L of HA solution with the concentration of around 7 mg/L and 13 mg/L for the artificial solar light and UV irradiations, respectively.

Before irradiation the preliminary adsorption of HA on the samples surface was determined. From the literature it is known that the adsorption equilibrium of HA on the TiO₂ surface is reached within 30 min^{5,7,8}. From our previous experiments it was estimated that this equilibrium is reached after around 3 hours⁴.

The solutions were first magnetically stirred in the dark for 3 hours and then were irradiated by 10 and 20 hours under UV and artificial solar light irradiation, respectively.

The changing of TOC in the reaction solution within the irradiation time was measured in the TOC analyser Multi N/C 3100 (Analytik Jena). The TOC analyses were performed for the monitoring of the mineralisation degree of HA after photocatalysis.

The plot of relative TOC changes in the HA solution against irradiation time was approximated to be linear, therefore the rate constant (k_{HA}), i.e., the slope of the linear relation, was determined on each sample and used as a measure of its photocatalytic activity. K_{HA} was determined as zero order kinetic reaction from equation: $C_{eq}-C = k_{HA} t$

where:

k_{HA} – rate constant of HA mineralisation

C_{eq} – concentration of HA (after adsorption on the photocatalyst surface)

C – concentration of HA after irradiation

RESULTS

XRD measurements

In Fig. 2 XRD patterns of TiO₂ original, calcined and modified with WO₃ are presented.

The original TiO₂ showed a poorly crystallized anatase phase with an insignificant amount of rutile. After calcination, the growing of anatase crystals could be observed. That resulted in the narrowing of the peaks, which refers to the anatase phase. Higher calcination temperature gives more sharp anatase reflexes. After the modification of TiO₂ with WO₃ there were not ob-

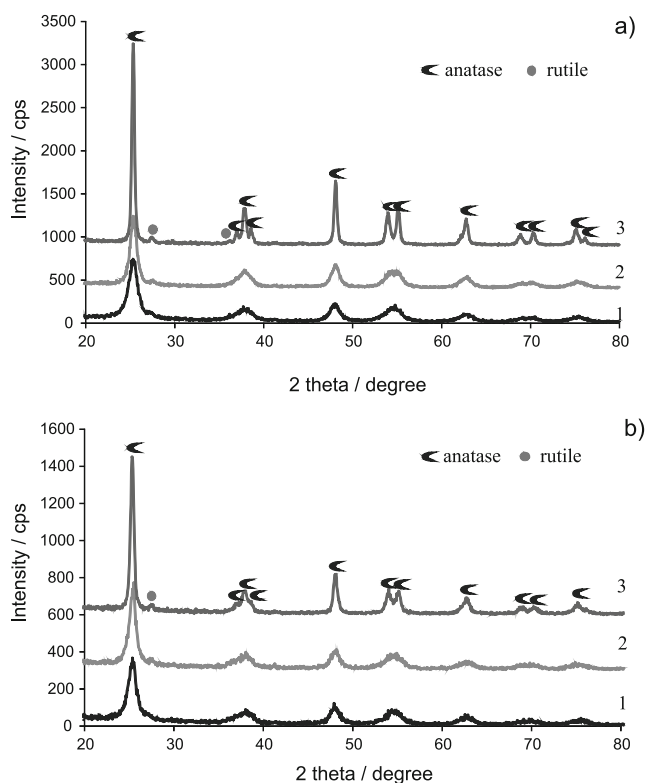


Figure 2. The XRD patterns of a) TiO₂ and b) WO₃/TiO₂ photocatalysts: (1) uncalcined, (2) calcined at 400°C and (3) calcined at 600°C

served any essential changes in the XRD patterns, the tungsten oxides phases were not detected. It could be caused by a good dispersion of tungsten oxides in a bulk of TiO₂ and also a small content of WO₃ in TiO₂ (3 wt %) being out of XRD detection limit. Similar results were reported by Shifu et al.¹⁹. They concluded that the reflections corresponding to WO₃ were not detected because of its too low concentration, or the presence in the amorphous state or by their strong dispersion in a bulk phase.

In order to check that crystallization of WO₃ occurs at the preparation conditions, tungsten precursor – WO₂, was dissolved in 30% H₂O₂ and was treated in a vacuum evaporator without an addition of TiO₂ and then followed calcinations at both, 400 and 600°C. The XRD patterns of the formed WO₃ are shown in Fig. 3.

From the XRD patterns it can be observed that WO₂ in the H₂O₂ solution underwent insignificant transformation to WO₃ (anorthic phase), which appeared quantitatively

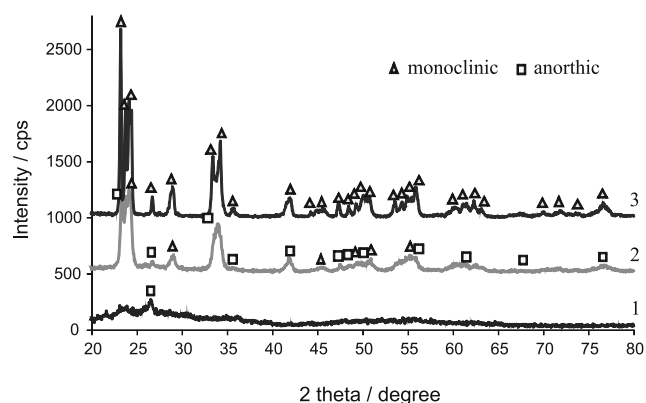


Figure 3. The XRD patterns of WO₂ dissolved in H₂O₂ and treated in vacuum evaporator, 1) uncalcined, 2) calcined at 400°C and 3) calcined at 600°C

after heat treatment at 400°C. At this temperature an insignificant amount of monoclinic phase of WO_3 was formed as well. At 600°C the only monoclinic WO_3 phase was observed. Observation of the crystal structure of the prepared WO_3/TiO_2 samples was performed by TEM analyses. In Fig. 4 TEM images of original TiO_2 and WO_3/TiO_2 calcined at 600°C are presented.

From the TEM images it is observed that TiO_2 consists of small particles, with the size below 10 nm, calcination at 600°C caused sintering and the growing of particles, but small anatase crystals of 5 nm size still remain in this WO_3/TiO_2 sample. The anatase crystals in WO_3/TiO_2 sample are surrounded with an amorphous state, which is most probably connected with doping of tungsten oxides to TiO_2 but also with the existence of some part of amorphous TiO_2 particles which did not transform to any crystal phase during the heat treatment.

UV-Vis /DR

In Fig. 5 the UV-Vis/DR spectra were shown for TiO_2 , WO_3/TiO_2 samples before and after calcination at 600°C and for comparison prepared at the same way WO_3 at 600°C

Modification of TiO_2 with WO_2 and H_2O_2 caused increase absorption of the visible light in the whole range but the most noticed in the range of 400–550 nm, similar as WO_3 . Heat treatment of WO_3/TiO_2 caused almost a complete reduction of absorption in the visible range. When we compare the spectra of TiO_2 and WO_3 , the latter one has shifted the absorption edge to the visible light and as a result its absorption maximum is at 457 nm whereas for TiO_2 is at 372 nm. It means that this WO_3 semiconductor can be excited with the visible light, contrary to TiO_2 and prepared WO_3/TiO_2 samples. In our previous studies when we used a higher amount of

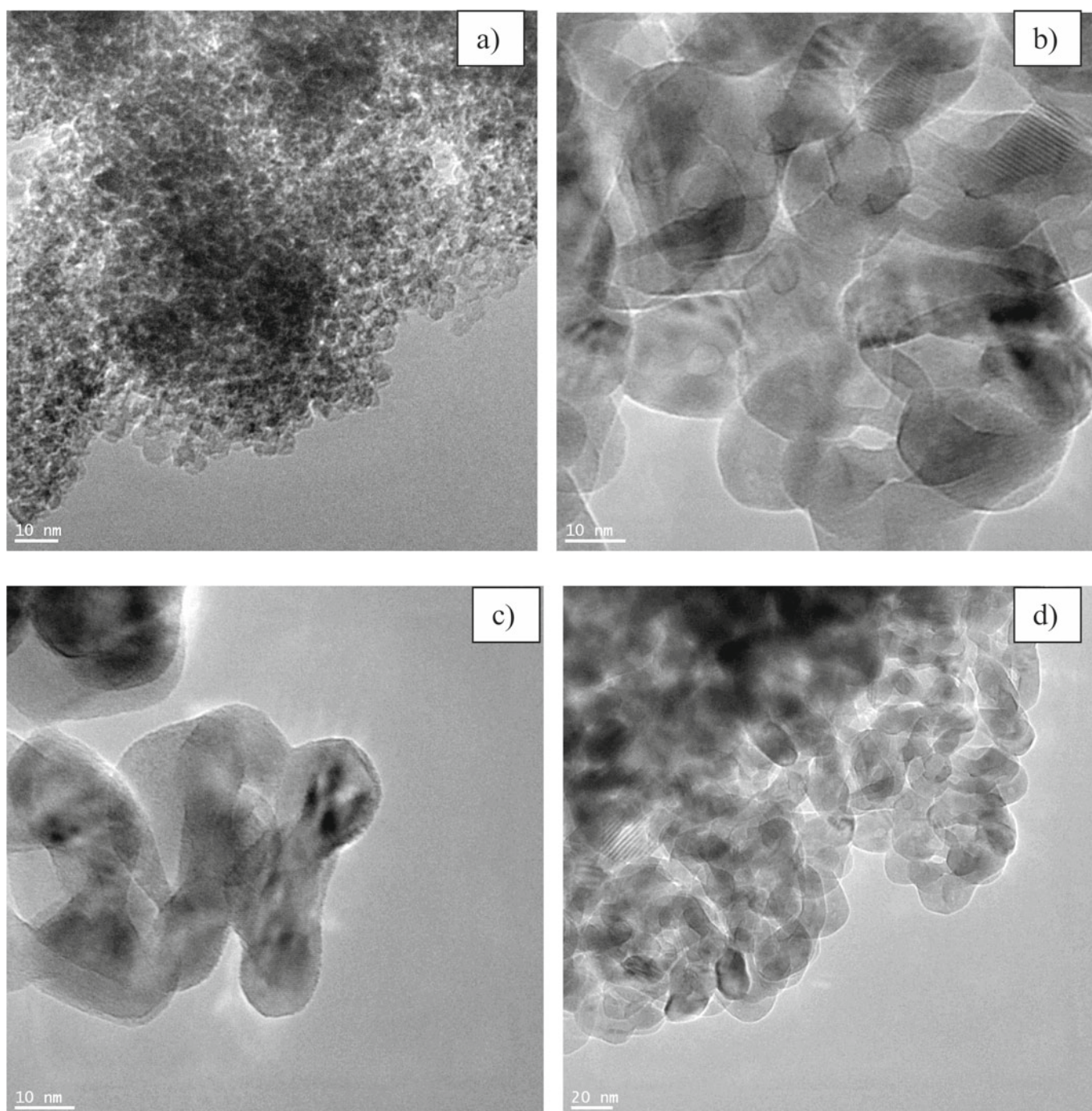


Figure 4. TEM images of a) TiO_2 , and b-d) WO_3/TiO_2 calcined at 600°C

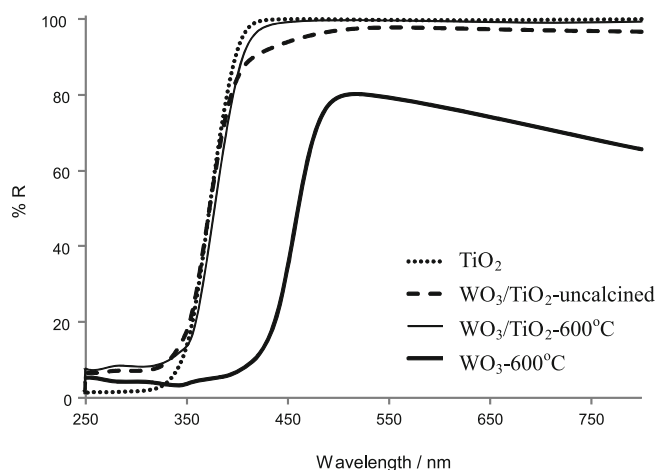


Figure 5. UV-Vis/DR spectra of TiO_2 , WO_3/TiO_2 and WO_3 samples

WO_3 doped to TiO_2 samples, they exhibited higher absorption of the visible light even after heat treatment, however their maximum of absorption was close to TiO_2 sample²².

OH radicals measurements

Detection of OH radicals formed on the photocatalysts surface after excitation with both, UV and artificial solar light irradiations was performed with the use of fluorescence method, i.e. the measurements of 7-hydroxycoumarin, the product of reaction coumarin with OH radicals. In Fig. 6 and 7 the plots showing the formation of 7-hydroxycoumarin within the time of irradiation are shown.

Modification of TiO_2 by tungsten precursor and H_2O_2 caused an increase of the OH radicals formation on the photocatalysts surface in comparison with the original TiO_2 under both, the artificial solar light and UV irradiation.

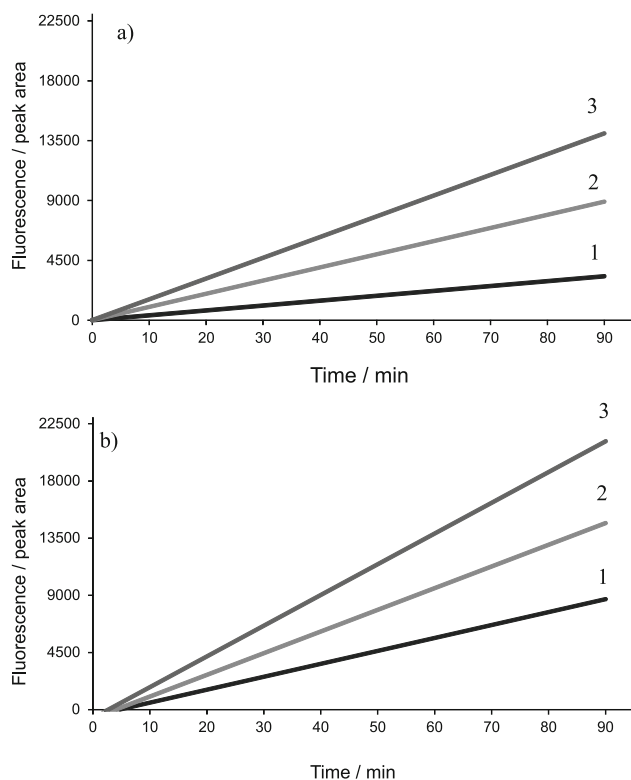


Figure 6. The OH radicals formation on: a) TiO_2 and b) WO_3/TiO_2 photocatalysts surface under UV light irradiation: (1) uncalcined, (2) calcined at 400°C and (3) calcined at 600°C

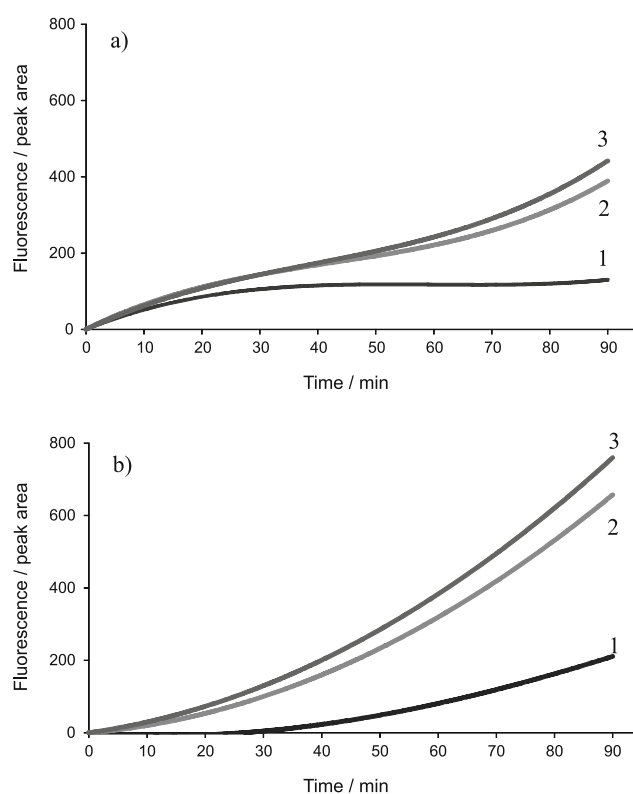


Figure 7. The OH radicals formation on: a) TiO_2 and b) WO_3/TiO_2 photocatalysts surface under the artificial solar light irradiation: (1) uncalcined, (2) calcined at 400°C and (3) calcined at 600°C

tions, which could be caused by a better separation of free carriers in the modified samples. The activity of the photocatalysts towards the OH radicals formation under the artificial solar light was incomparably less effective than under UV, it is assumed that this activity was connected just with the excitation of the photocatalysts by a small content of UV light emitted by the fluorescence lamp. With the increase of the heat treatment temperature to 600°C, higher activity towards the OH radicals formation was observed, which can be explained by the growing of anatase crystals, as it was already reported in our previous paper²². The linear correlation between the time of UV irradiation and the fluorescence peak of 7-hydroxycoumarin indicates that the amount of OH radicals is proportionally increasing to the irradiation time, under the solar light irradiation the concentration of OH radicals seems to be insignificant and therefore at the beginning the reaction with coumarin does not occur, with the time of irradiation the amount of OH radicals increases and then the product of reaction, 7-hydroxycoumarin appears, at the higher concentration of the OH radicals the trend of this line tends to be linear. Probably for the detection of a small amount of OH radicals the concentration of coumarin was too high. We suppose that for the coupled TiO_2 - WO_3 photocatalysts injecting of electrons from the conductive band of TiO_2 to WO_3 took place, making the TiO_2 particles more positively charged and improving the separation of free carriers, therefore with the time of irradiation the amount of the formed OH radicals increased rapidly. In Table 1 the values of k_{OH} coefficient for the OH radicals formation under the UV irradiation were listed.

Table 1. Chemical composition of the spent vanadium catalyst

Properties	Photocatalyst					
	TiO ₂			WO ₃ /TiO ₂		
	uncalcined	400°C	600°C	uncalcined	400°C	600°C
Zeta potential [mV]	12	8,6	9,8	-10,3	-15,4	-23,9
S _{BET} [m ² /g]	275	172	51	212,6	185,7	66,3
k _{OH} (UV) [min ⁻¹] (R ²)	31,62 (0,97)	94,65 (0,99)	154,15 (0,99)	101,72 (0,98)	170,79 (0,99)	241,91 (0,99)
k _{HA} (UV) [mg/dm ³ h] (R ²)	0,83 (0,99)	0,63 (0,99)	1,51 (0,99)	1,58 (0,99)	0,89 (0,99)	1,12 (0,99)

Zeta Potential

The values of electrokinetic zeta potentials of the photocatalysts particles measured in the ultra pure water are listed in Table 1.

Modification of TiO₂ by tungsten oxides caused the change of the zeta potential of particles from +12 to -10.3 mV. With an increase of the temperature of the heat treatment the changes were higher in direct minus values. It means that these TiO₂ photocatalysts after a modification with tungsten oxides exhibited the Bronstead acid behaviour. A similar effect was observed by the other authors²⁶.

Measurements of BET surface area

The values of the SBET surface area for the studied photocatalysts are listed in Table 1.

Modification of TiO₂ by tungsten oxides caused that

the SBET surface area decreased from 275 to 212.6 m²/g. Calcination of the samples caused the reduction of the BET surface area due to the growth and sintering of the particles. However, the BET surface area of TiO₂ heated at 600°C of around 50 m²/g was comparable with the commercially available TiO₂ P25 of Evonik company.

Photocatalytic decomposition of HA

HA were initially adsorbed on the photocatalyst surface for 3 h (in a dark) and then were irradiated with UV either artificial solar light irradiations. In Fig. 8 and 9 the changes of the TOC concentration after adsorption and during the UV and artificial solar light irradiations for TiO₂ and the WO₃/TiO₂ samples are presented.

The original TiO₂ showed high adsorption of HA on its surface, around 47.15 mg HA/g of photocatalyst. A

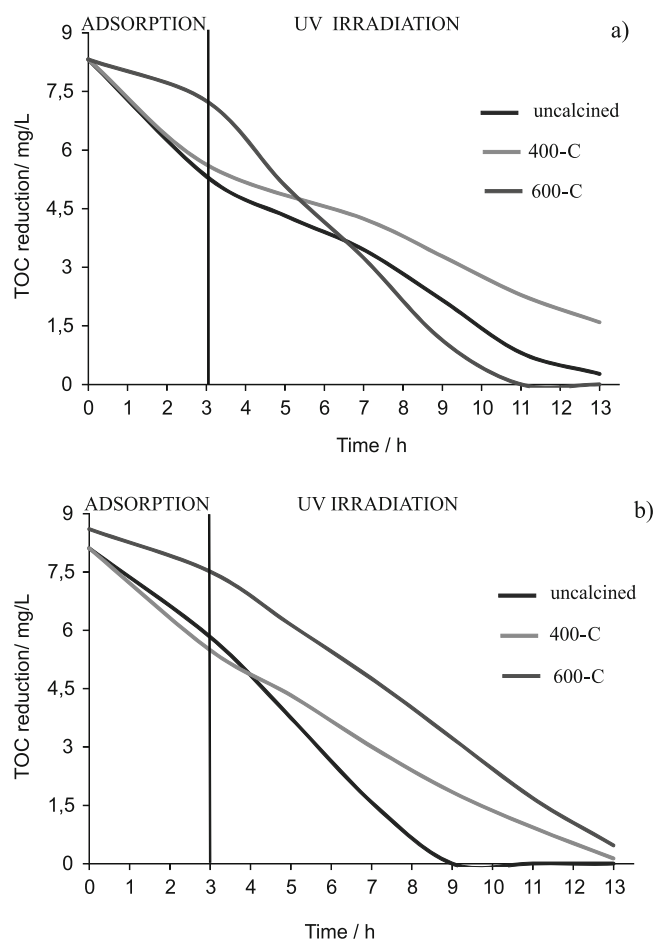


Figure 8. Adsorption and photodecomposition of HA under UV irradiation on a) TiO₂ and b) WO₃/TiO₂ photocatalysts

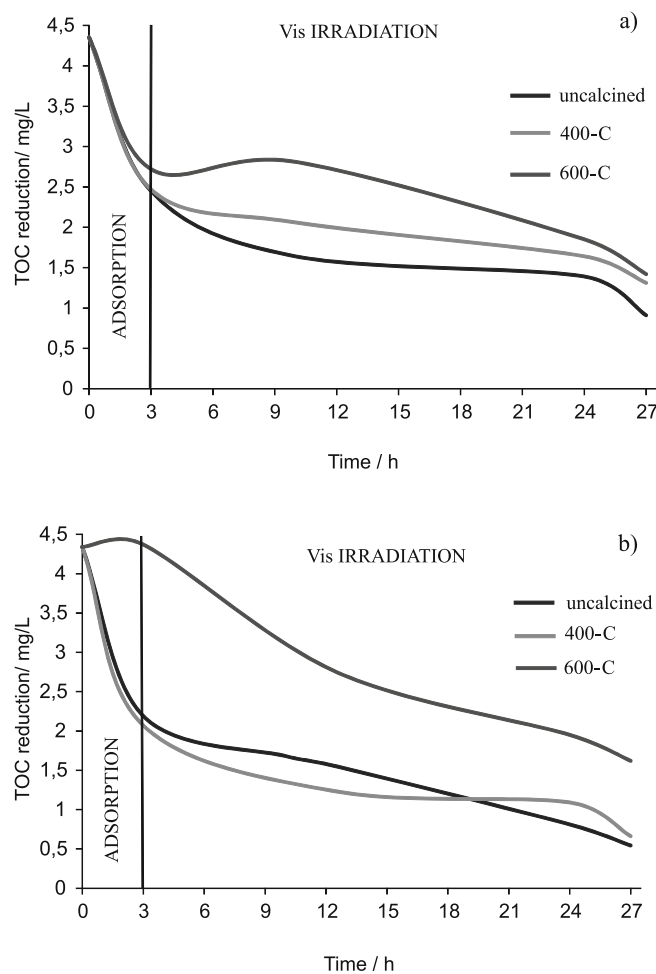


Figure 9. Adsorption and photodecomposition of HA under artificial solar light irradiation on a) TiO₂ and b) WO₃/TiO₂ photocatalysts

similar value of HA adsorption on TiO₂ was obtained by Dziejczak et al.⁸, whereas the other researchers reported the lower values, from 0.7 to 23 mg/g photocatalyst^{5,7}. It was reported that adsorption of HA is favored in acid conditions and decreases with the increasing of the pH of the solution⁶⁻⁸. In our experiment we used the natural pH of HA solution, which was around 6.6. Doping WO₃ to TiO₂ caused that the adsorption of HA was insignificant lower, around 35.9 mg HA/g of photocatalyst. Calcination of the samples led to a decrease of HA adsorption, what was related to their lower BET surface area in comparison with the uncalcined ones. In general adsorption of HA on the surface of both, TiO₂ and WO₃/TiO₂ photocatalysts was strongly related to their BET surface area, changing of the electrokinetic potential in the WO₃/TiO₂ samples did not influence the HA affinity to their surface. The photocatalytic decomposition of HA was performed after its adsorption on the photocatalyst surface, therefore to evaluate the photocatalytic activity of the samples, the concentration of HA after adsorption was taken as an initial concentration of HA. The rate of HA decomposition under UV irradiation was calculated and the values of k_{HA} were listed in Table 1. Under artificial solar light irradiation it was difficult to fit the kinetic curve, because the mineralisation degree was not increasing proportionally with the irradiation time. It could be caused by the slowly increasing of the OH radicals during the excitation of the photocatalysts with the visible light. It was reported that the process of HA decomposition does not follow the Langmuir – Hisnshelwood kinetics and is controlled rather by the rate of the generation of the reactive radicals in the photoexcitation process⁸. In our results in the case of TiO₂ there is almost the same relation, the higher amount of the OH radicals formation on TiO₂ calcined at 600°C resulted in a higher mineralisation degree of HA in comparison to uncalcined one TiO₂. In the case of the WO₃/TiO₂ samples there is not the same relation, the uncalcined WO₃/TiO₂ showed the highest mineralisation degree, however this WO₃/TiO₂ calcined at 600°C (with the highest amount of OH radicals formation) was more active than that calcined at 400°C. Therefore it can be concluded that in the case of the WO₃/TiO₂ sample the surface oxidative reactions are very important. Cho and Choi⁶ reported that since HA is a macromolecule with many redox centers, a series of electron transfers may take place on TiO₂ surface. Although the photoinduced electron-donating (oxidation) property of HA has been mainly discussed so far, the electron-accepting (reduction) property of HA should be also recognized in order to understand the overall process. Sequential electron transfers from excited HA to TiO₂ conductive band could lead to mineralization with CO₂ evolution while the electron transfers from conductive band to HA tend to inhibit the mineralization process⁶. In the case of WO₃/TiO₂ sample there is a possibility of electron transfer from excited HA to both TiO₂ and WO₃, and it seems that the transfer to WO₃ plays an important role in HA mineralisation.

To summarize the photocatalytic activity of TiO₂ and WO₃/TiO₂ photocatalysts, it has to be pointed out that in the case of TiO₂ enhanced photocatalytic activity was observed for the sample calcined at 600°C which revealed

high OH radicals formation. However, in the case of WO₃/TiO₂ photocatalysts the highest photocatalytic activity was achieved on the uncalcined sample, which showed high adsorption of HA, that could be caused by a better separation of free carriers in TiO₂ and enhanced the yield of surface oxidative reactions. Self-excitation of HA preliminary adsorbed on the surface of the photocatalyst could occur resulting in their fast decomposition, especially under UV irradiation. The photocatalytic activity of the TiO₂ and WO₃/TiO₂ photocatalysts under artificial solar light was rather poor, and can be assigned to the small content of UV in the incident light and also in the case of WO₃/TiO₂ photocatalysts, the presence of WO₃, which can be excited with the visible light. In that case WO₃/TiO₂ sample calcined at 600°C was the most active due to the highest amount of OH radicals formation in comparison with the other samples.

CONCLUSIONS

Modification of TiO₂ by WO₃ caused better separation of free carriers which resulted in both, the increased OH radicals formation on the photocatalyst surface and enhanced direct oxidation of HA. Calcination of TiO₂ and WO₃/TiO₂ photocatalysts at 400 and 600°C caused the increased OH radicals formation but also the reduction of BET surface area. Mineralisation of HA on TiO₂ was dependent on OH radicals formation, being the highest for TiO₂ calcined at 600°C, which revealed the highest OH radicals formation. However, in the case of the WO₃/TiO₂ photocatalysts high adsorption of HA on the photocatalyst surface facilitated their decomposition under UV irradiation, and the highest decomposition rate of HA was noted for the uncalcined sample. In that case a good separation of free carriers and direct oxidation of HA occurred, which surely accelerated their decomposition. Possible self-excitation of HA under UV irradiation could enhance their decomposition. The prepared WO₃/TiO₂ photocatalysts did not show narrowing of the band gap and the red shift in the UV/Vis absorption spectra, therefore they were not active under the visible light. The decomposition of HA under the artificial solar light irradiation on WO₃/TiO₂ photocatalyst calcined at 600°C can be assigned to the small content of UV in the incident light and higher OH radicals formation in comparison with the other samples.

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