

Acoustical studies on molecular interaction of 1,3,4-pyrazoline derivatives using ultrasonic technique at 303.15 K

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ABSTRACT

The acoustical parameters for the binary liquid mixtures containing Pyrazoline derivatives in DMF have been estimated at 303.15 K, from the measured values of ultrasonic velocity (U), density (ρ) and viscosity (η). From these data some of acoustical parameters such as adiabatic compressibility (β), free length (L_f), acoustic impedance (Z), Rao's constant (R), molar compressibility (W), relaxation time (τ), free volume (V_f) and internal pressure (π_i), etc., have been computed using the standard relations. The results have been discussed in terms of molecular interactions.

Keywords: Binary mixture; molecular interaction; ultrasonic velocity; density; viscosity; free length

1. INTRODUCTION

The study of ultrasonics in organic liquids has been the subject of extensive research recently. Ultrasonic methods have the added advantage of being less cost with efficiency comparable to other methods. Hence, a number of works have reported the study through ultrasonic method [1-4]. Acoustics is a field widely used in recent years to study various molecular interactions. The study of the liquid state properties using spectroscopic and acoustical methods provides valuable information for their varied usage.

The study of liquid mixtures containing of polar and non-polar components find applications in industrial and technological process [5]. The mixing of different components gives rise to solutions that generally do not behave ideally [6-7]. Further these properties have been widely used to study the molecular interaction between the various species in the mixture [8-9]. In recent years, the measurement of ultrasonic velocity has been extensively applied in understanding the nature of molecular systems, physicochemical behaviour and molecular interactions in liquid mixtures [10-12].

Intermolecular interaction in various binary liquid mixtures at different temperatures have been studied by several authors[13-16]. Ultrasonic velocity and related thermodynamic parameters helps us for characterizing thermodynamic and physico-chemical aspects of binary liquid mixtures such as molecular association and dissociation [17-18]. Viscosity, density

measurements and the properties derived from these are excellent tools to detect solute – solute and solute – solvent interactions.

Physico-chemical properties like density, viscosity and speed of sound have got considerable importance in forming theoretical models as well as their applications in a number of branches of science. A considerable progress has been made theoretical understanding of liquid-liquid mixture [19-22].

2. EXPERIMENTAL

2. 1. Choice of Solvent

Dimethyl sulphoxide is chosen as solvent in the present work. This solvent is of industrial interest because of their wide use as solvent and solubilizing agent.

The densities, viscosities and ultrasonic velocities of solvent and solutions of different concentrations were measured at 303.15 K by specific gravity bottle, an Ostwald's viscometer and single frequency ultrasonic interferometer operating at 2 MHz.

2. 2. Density

The density of pure liquids and mixtures are measured using a 10ml specific gravity bottle. The measured density was calculated using the formula

$$\rho_2 = (w_2/w_1)\rho_1$$

where:

w_1 is the weight of the distilled water

w_2 is the weight of the experimental liquid

ρ_1 is the density of water

2. 3. Viscosity

The viscosity of the pure liquids and liquid mixtures are measured using an Ostwald's Viscometer calibrated with doubly distilled water. The Ostwald's Viscometer with the experimental liquid is immersed in a temperature controlled water bath at 303.15 K. The digital stopwatch, with an accuracy of ± 0.01 sec was used to determine flow time of solutions. Using the flow times (t) and known of standard water sample, the viscosity of solvent and solutions were determined according the following equation:

$$\eta_2 = \eta_1(t_2/t_1)(\rho_2/\rho_1)$$

where:

η_1 is the viscosity of water

t_1 is the time of flow of water

ρ_1 is the density of water

η_2 is the viscosity of the binary mixture

t_2 is the time of flow of the binary mixture

ρ_2 is the density of the binary mixture

2. 4. Ultrasonic velocity

The sound velocity of the liquid mixture have been measured using an ultrasonic interferometer (Mittal Enterprises, New Delhi) working at a fixed frequency of 2 MHz. The binary liquid mixture is filled in the measuring cell with quartz crystal and then micrometer was fixed. The circulation of water from the thermostat at 303.15 K was started and test solution in the cell is allowed to thermally equilibrate. The micrometer was rotated very slowly so as to obtain a maximum or minimum of anode current (n). A number of maximum reading of anode current were counted. The total distance (d) travel by the micrometer for n = 10, was read. The wave length (λ) was determined according to the following equation:

$$\lambda = 2d/n$$

The sound velocity (U) of solvent and solutions were calculated from the wavelength and frequency (F) according to the following equation:

$$U = \lambda F$$

3. THEORY

The ultrasonic velocity (U), density (ρ) and viscosity (η), molar volume (Vm) in pure liquids and liquid mixtures of various concentrations have been measured at 303.15K. Thermodynamic and acoustical parameters such as adiabatic compressibility (β), free length (L_f), acoustic impedance (Z), relative association (R_A), relaxation strength (r), Rao's constant (R), molar compressibility (W), relaxation time (τ), Vander waal's constant (b), Isothermal compressibility (β_T), Isothermal expansion coefficient (α), ultrasonic attenuation (α/f^2), free volume (V_f) and internal pressure (π_i) were determined using the observed values of velocity, density & viscosity using the standard relations given below.

1. The adiabatic compressibility (β) has been calculated from sound velocity (U) and the density (ρ) of the medium using the relation

$$\beta = 1/U^2\rho \quad \text{--- (1)}$$

2. Intermolecular free length (L_f) is calculated using the standard expression

$$L_f = K\beta^{1/2} \quad \text{--- (2)}$$

where K is a Jacobson's constant (= 2.0965×10^{-6})

3. Acoustic impedance (Z) was calculated by the equation

$$Z = U\rho \quad \text{--- (3)}$$

where ρ is the density of the mixture and U is the ultrasonic velocity of the mixture.

4. Molar compressibility or Wada's constant (W) can be calculated by the following equation

$$W = (M/\rho)\beta^{-1/7} \quad \text{--- (4)}$$

where M is the molecular weight of the solution which can be calculated according to

the equation

$$M = M_1X_1 + M_2X_2 \quad \text{--- (5)}$$

where X_1 and X_2 are mole fractions of solvent and solute, respectively. M_1 and M_2 are the molecular weights of the solvent and solute respectively.

5. The molar sound velocity or Rao's constant (R) was calculated by the equation

$$R = (M/\rho)U^{1/3} \quad \text{---(6)}$$

where ρ is the density and U is the ultrasonic velocity of the mixture.

6. The Relative association (R_A) was calculated by the following equation

$$R_A = \rho/\rho_o(U_o/U)^{1/3} \quad \text{---(7)}$$

where U , U_o and ρ , ρ_o are ultrasonic velocities and densities of solution and solvent respectively.

7. Relaxation strength (r) was calculated by the following equation

$$r = 1-[U/U_\infty]^2 \quad \text{---(8)}$$

where $U_\infty = 1600$ m/sec.

8. Vander Waals constant (b) was calculated by the following equation

$$b = (M/\rho)[1-(RT/MU^2)\{(1+MU^2/3RT)^{1/2} - 1\}] \quad \text{---(9)}$$

where R is the gas constant ($=8.314 \text{ JK}^{-1} \text{ mol}^{-1}$) and T is the absolute temperature.

9. Viscous relaxation time (τ) was calculated by the following equation

$$\tau = 4\eta/3\rho U^2 \quad \text{--- (10)}$$

where η is the viscosity, ρ is the density and U is the ultrasonic velocity of the mixture.

10. Isothermal compressibility, β_T

$$\beta_T = 1.71 \times 10^{-3}/T^{4/9}U^2\rho^{4/3} \quad \text{---(11)}$$

where, U is the ultrasonic velocity, ρ is the density and T is the Temperature.

11. Isothermal Expansion coefficient, α

$$\alpha = (0.0191 \beta_T)^{1/4} \quad \text{---(12)}$$

12. Ultrasonic attenuation (α/f^2)

$$(\alpha/f^2) = 8\pi^2\eta/3\rho U^2 \quad \text{--- (13)}$$

13. Free volume (V_f) was calculated by the following equation

$$V_f = [MU/K\eta]^{3/2} \quad \text{---(14)}$$

where K is a constant ($= 4.28 \times 10^9$)

14. The Internal pressure (π_i) was calculated by the following equation

$$\pi_i = bRT [K\eta/U]^{1/2} \rho^{2/3} / M^{7/6} \quad \text{---(15)}$$

where b is the packing factor (=2), K is a constant (= 4.28 X 10⁹)

The experimental density (ρ), viscosity (η) ultrasonic velocity (U) for Pyrazoline derivatives at various concentrations are given in Tables (1-3). The computed acoustical parameters for the above Pyrazoline derivatives are given in Tables (4-12).

Table 1. Ultrasonic velocity (U), Density(ρ), Viscosity (η) and Molar volume (Vm) of Pyrazoline derivatives with DMF at 303.15 K.

Concentration	Pyrazoline derivatives (P1-P10)			
	U	ρ	η	Vm
(M)	(m/s)	(Kg/m ³)	(mPa·s)	(10 ⁻⁶ m ³ /mole)
Compound P1				
0.02	1457	961.7	0.9288	83.98
0.04	1470	978.2	0.9517	90.41
0.06	1480	997.2	0.9702	96.39
0.08	1490	1012.5	0.9923	102.52
0.1	1496	1032.8	1.017	107.94
Compound P2				
0.02	1487	961.4	0.956	84.09
0.04	1493	978	0.9725	90.59
0.06	1495	997.1	0.9917	96.64
0.08	1508	1000.3	1.009	104.09
0.1	1512	1010.8	1.0135	110.68
Compound P3				
0.02	1437	961.8	0.9771	85.09
0.04	1440	978.3	0.9938	92.61
0.06	1447	980.7	1.0033	101.32
0.08	1509	996	1.0332	108.55
0.1	1511	1016.3	1.0542	115.01

Table 2. Ultrasonic velocity (U), Density (ρ), Viscosity (η) and Molar volume (Vm) of Pyrazoline derivatives with DMF at 303.15 K.

Concentration	Pyrazoline derivatives (P1-P10)			
	U	ρ	η	Vm
(M)	(m/s)	(Kg/m ³)	(mPa·s)	(10 ⁻⁶ m ³ /mole)

Compound P4				
0.02	1464	961.3	0.9697	84.62
0.04	1469	979.1	0.9876	91.51
0.06	1475	980.6	0.9962	99.80
0.08	1498	995.6	1.0127	106.59
0.1	1504	1015.7	1.0391	112.61
Compound P5				
0.02	1434	962	0.9222	83.79
0.04	1457	981.7	0.9411	89.76
0.06	1474	1000.8	0.9737	95.56
0.08	1486	1019.3	0.999	101.20
0.1	1515	1037.2	1.0166	106.71
Compound P6				
0.02	1406	963.4	0.9235	83.27
0.04	1464	981.4	0.9478	89.02
0.06	1478	1000.4	0.9733	94.46
0.08	1481	1018.8	0.9912	99.76
0.1	1517	1037	1.0238	104.89

Table 3. Ultrasonic velocity (U), Density (ρ), Viscosity (η) and Molar volume (V_m) of Pyrazoline derivatives with DMF at 303.15 K.

Concentration	Pyrazoline derivatives (P1-P10)			
	U	ρ	η	V_m
(M)	(m/s)	(Kg/m ³)	(mPa·s)	(10 ⁻⁶ m ³ /mole)
Compound P7				
0.02	1430	963.1	0.9646	84.00
0.04	1452	980.8	0.9823	90.46
0.06	1472	999.9	1.0086	96.55
0.08	1486	1019.8	1.0433	102.33
0.1	1512	1037.9	1.0692	108.08
Compound P8				
0.02	1413	961.2	0.9367	83.84
0.04	1414	976.1	0.9457	90.24
0.06	1430	995.2	0.9836	96.04
0.08	1435	1000.3	0.994	103.05
0.1	1446	1019.2	1.0099	108.49
Compound P9				

0.02	1403	962.7	0.9345	83.62
0.04	1420	975.1	0.9508	90.17
0.06	1446	989.1	0.9672	96.39
0.08	1469	999.6	0.9802	102.80
0.1	1480	1008.3	1.0014	109.27
Compound P10				
0.02	1448	967.6	0.9494	83.53
0.04	1473	984.5	0.9727	89.96
0.06	1476	994.1	0.9816	96.87
0.08	1477	1007.5	0.9958	103.26
0.1	1486	1014.3	1.0121	110.20

Table 4. Specific Acoustic Impedance (Z), Adiabatic compressibility (β_{ad}), Intermolecular Free Length (L_f), Rao's Constant (R) and Molar compressibility (W) of Pyrazoline derivatives with DMF at 303.15 K.

Concentration	Pyrazoline derivatives (P1-P10)				
	Z	β_{ad}	L_f	R	W
(M)	($10^3 \text{Kg/m}^2\text{s}$)	(10^{-10}Pa^{-1})	(10^{-10}m)	(m^3/mole) (m/s) ^{1/3}	(m^3/mole) (N/m^2) ^{1/7}
Compound P1					
0.02	1401.197	4.8983	0.4639	0.9521	1.7955
0.04	1437.954	4.7308	0.4559	1.0280	1.9426
0.06	1475.856	4.5782	0.4485	1.0985	2.0808
0.08	1508.625	4.4487	0.4421	1.1709	2.2221
0.1	1545.069	4.3263	0.4360	1.2345	2.3489
Compound P2					
0.02	1429.602	4.7041	0.4547	0.9598	1.8083
0.04	1460.154	4.5871	0.4490	1.0354	1.9552
0.06	1490.665	4.4872	0.4441	1.1050	2.0922
0.08	1508.452	4.3961	0.4395	1.1936	2.260
0.1	1528.33	4.3274	0.4361	1.2704	2.4086
Compound P3					
0.02	1382.107	5.035	0.4704	0.9603	1.8122
0.04	1408.752	4.9295	0.4654	1.0458	1.9782
0.06	1419.073	4.87	0.4626	1.1460	2.1679
0.08	1502.964	4.4092	0.4402	1.2451	2.3560
0.1	1535.629	4.3097	0.4352	1.3197	2.5041

Table 5. Specific Acoustic Impedance (Z), Adiabatic compressibility (β_{ad}), Intermolecular Free Length (L_f), Rao's Constant (R) and Molar compressibility (W) of Pyrazoline derivatives with DMF at 303.15 K.

Concentration	Pyrazoline derivatives (P1-P10)				
	Z	β_{ad}	L_f	R	W
(M)	($10^3 \text{Kg/m}^2\text{s}$)	(10^{-10}Pa^{-1})	(10^{-10}m)	(m^3/mole) (m/s) ^{1/3}	(m^3/mole) (N/m^2) ^{1/7}
Compound P4					
0.02	1407.343	4.8535	0.4618	0.9608	1.8115
0.04	1438.298	4.7329	0.4560	1.0403	1.9662
0.06	1446.385	4.6873	0.4538	1.1360	2.1471
0.08	1491.409	4.476	0.4435	1.2196	2.3084
0.1	1527.613	4.3525	0.4373	1.2902	2.4485
Compound P5					
0.02	1379.508	5.0551	0.4713	0.9449	1.7833
0.04	1430.337	4.7985	0.4592	1.0176	1.9248
0.06	1475.179	4.5989	0.4495	1.0876	2.0616
0.08	1514.68	4.4428	0.4419	1.1549	2.1941
0.1	1571.358	4.2006	0.4296	1.2255	2.3320
Compound P6					
0.02	1354.54	5.2508	0.4804	0.9329	1.7628
0.04	1436.77	4.7541	0.4571	1.0108	1.9113
0.06	1478.591	4.5759	0.4484	1.0760	2.0393
0.08	1508.843	4.4751	0.4435	1.1372	2.1606
0.1	1573.129	4.1903	0.4291	1.2053	2.2932

Table 6. Specific Acoustic Impedance (Z), Adiabatic compressibility (β_{ad}), Intermolecular Free Length (L_f), Rao's Constant (R) and Molar compressibility (W) of Pyrazoline derivatives with DMF at 303.15 K.

Concentration	Pyrazoline derivatives (P1-P10)				
	Z	β_{ad}	L_f	R	W
(M)	($10^3 \text{Kg/m}^2\text{s}$)	(10^{-10}Pa^{-1})	(10^{-10}m)	(m^3/mole) (m/s) ^{1/3}	(m^3/mole) (N/m^2) ^{1/7}
Compound P7					
0.02	1377.233	5.0776	0.4724	0.9464	1.7868
0.04	1424.122	4.836	0.4610	1.0243	1.9375
0.06	1471.853	4.6156	0.4504	1.0983	2.0818
0.08	1515.423	4.4407	0.4417	1.1678	2.2187
0.1	1569.305	4.2144	0.4303	1.2405	2.3609
Compound P8					
0.02	1358.176	5.2108	0.4785	0.9408	1.7767
0.04	1380.205	5.124	0.4745	1.0128	1.9169

0.06	1423.136	4.9138	0.4647	1.0820	2.0524
0.08	1435.431	4.8547	0.4619	1.1623	2.2059
0.1	1473.763	4.6925	0.4541	1.2269	2.3338
Compound P9					
0.02	1350.668	5.2771	0.4816	0.9361	1.7689
0.04	1384.642	5.086	0.4728	1.0135	1.9174
0.06	1430.239	4.8353	0.4610	1.0900	2.0646
0.08	1468.412	4.6359	0.4513	1.1686	2.2152
0.1	1492.284	4.5278	0.4461	1.2452	2.3626
Compound P10					
0.02	1401.085	4.9291	0.4654	0.9450	1.7843
0.04	1450.169	4.6814	0.4536	1.0235	1.9357
0.06	1467.292	4.6174	0.4504	1.1030	2.0886
0.08	1488.078	4.5498	0.4471	1.1760	2.2312
0.1	1507.25	4.4647	0.4429	1.2575	2.3875

Table 7. Relaxation time(τ), van der Waals constant (b), Relaxation strength (r), Relative association (R_A) and Ultrasonic attenuation (α/f^2) of Pyrazoline derivatives with DMF at 303.15 K.

Concentration (M)	Pyrazoline derivatives (P1-P10)				
	τ (10^{-10} sec)	b ($m^3/mole$)	r	R_A	α/f^2 (10^{-12} sec)
Compound P1					
0.02	6.066	0.0838	0.1707	1.0215	8.2097
0.04	6.0031	0.0902	0.1558	1.0360	8.0528
0.06	5.9224	0.0962	0.1443	1.0537	7.8908
0.08	5.8859	0.1023	0.1327	1.0675	7.7896
0.1	5.8665	0.1077	0.1257	1.0874	7.7328
Compound P2					
0.02	5.9961	0.0839	0.1362	1.0143	7.9515
0.04	5.948	0.0904	0.1292	1.0304	7.8559
0.06	5.9333	0.0964	0.1269	1.0501	7.8261
0.08	5.9142	0.1038	0.1116	1.0504	7.7336
0.1	5.8478	0.1104	0.1069	1.0605	7.6266
Compound P3					
0.02	6.5596	0.0849	0.1933	1.0264	9.0014
0.04	6.5319	0.0924	0.19	1.0432	8.9447
0.06	6.5147	0.1011	0.1821	1.0441	8.8780
0.08	6.0741	0.1083	0.1105	1.0457	7.9375
0.1	6.0577	0.1148	0.1081	1.0665	7.9056

Table 8. Relaxation time(τ), van der Waals constant (b), Relaxation strength (r), Relative association (R_A) and Ultrasonic attenuation (α/f^2) of Pyrazoline derivatives with DMF at 303.15 K.

Concentration	Pyrazoline derivatives (P1-P10)				
	τ	b	r	R_A	α/f^2
(M)	(10^{-10} sec)	($m^3/mole$)			(10^{-12} sec)
Compound P4					
0.02	6.2753	0.0844	0.1627	1.0195	8.4524
0.04	6.2323	0.0913	0.1570	1.0372	8.3659
0.06	6.226	0.0996	0.1501	1.0374	8.3235
0.08	6.0438	0.1063	0.1234	1.0478	7.9558
0.1	6.0302	0.1124	0.1164	1.0675	7.9063
Compound P5					
0.02	6.2157	0.0836	0.1967	1.0273	8.5473
0.04	6.0211	0.0895	0.1707	1.0428	8.1490
0.06	5.9707	0.0953	0.1512	1.0590	7.9875
0.08	5.9179	0.1010	0.1374	1.0756	7.8529
0.1	5.6938	0.1065	0.1034	1.0875	7.4110
Compound P6					
0.02	6.4654	0.0831	0.2277	1.0356	9.0678
0.04	6.008	0.0888	0.1627	1.0408	8.0923
0.06	5.9383	0.0942	0.1466	1.0576	7.9227
0.08	5.9143	0.0995	0.1432	1.0763	7.8747
0.1	5.7201	0.1047	0.1011	1.0868	7.4354

Table 9. Relaxation time(τ), van der Waals constant (b), Relaxation strength (r), Relative association (R_A) and Ultrasonic attenuation (α/f^2) of Pyrazoline derivatives with DMF at 303.15 K.

Concentration	Pyrazoline derivatives (P1-P10)				
	τ	b	r	R_A	α/f^2
(M)	(10^{-10} sec)	($m^3/mole$)			(10^{-12} sec)
Compound P7					
0.02	6.5304	0.0838	0.2012	1.0294	9.0052
0.04	6.3339	0.0902	0.1764	1.0430	8.6018
0.06	6.2071	0.0963	0.1536	1.0585	8.3150
0.08	6.1773	0.1021	0.1374	1.0762	8.1972
0.1	6.0081	0.1078	0.1069	1.0889	7.8356
Compound P8					
0.02	6.5079	0.0837	0.2201	1.0315	9.0821

0.04	6.461	0.0900	0.2189	1.0472	9.0102
0.06	6.4443	0.0958	0.2012	1.0637	8.8864
0.08	6.4341	0.1028	0.1956	1.0679	8.8415
0.1	6.3186	0.1082	0.1832	1.0853	8.6167
Compound P9					
0.02	6.5752	0.0834	0.2311	1.0355	9.2415
0.04	6.4477	0.0899	0.2123	1.0447	8.9537
0.06	6.2356	0.0961	0.1832	1.0533	8.5035
0.08	6.0588	0.1026	0.1570	1.0589	8.1330
0.1	6.0455	0.1091	0.1443	1.0655	8.0549
Compound P10					
0.02	6.2396	0.0833	0.1809	1.0299	8.4972
0.04	6.0715	0.0897	0.1524	1.042	8.1279
0.06	6.0432	0.0966	0.1489	1.0514	8.0737
0.08	6.0409	0.1031	0.1478	1.0653	8.0651
0.1	6.025	0.1100	0.1374	1.0704	7.9951

Table 10. Isothermal compressibility (β_T), Isothermal expansion co-efficient (α), Free volume (V_f) and Internal pressure (π_i) of Pyrazoline derivatives with DMF at 303.15 K

Concentration (M)	Pyrazoline derivatives (P1-P10)			
	β_T ($10^{-14} \text{m}^2 \text{N}^{-1}$)	α (10^4K^{-1})	V_f ($10^{-6} \text{m}^3/\text{mole}$)	π_i (10^6Pa)
Compound P1				
0.02	0.6694	1.0633	0.1610	4.8309
0.04	0.6429	1.0526	0.1803	4.4289
0.06	0.6182	1.0424	0.2005	4.0963
0.08	0.5976	1.0336	0.2197	3.8133
0.1	0.5774	1.0247	0.2371	3.5921
Compound P2				
0.02	0.6429	1.0527	0.1592	4.8448
0.04	0.6234	1.0446	0.1791	4.4325
0.06	0.6059	1.0372	0.1977	4.1084
0.08	0.5930	1.0316	0.2192	3.7778
0.1	0.5817	1.0266	0.2435	3.5014
Compound P3				
0.02	0.6881	1.0707	0.1491	4.9129
0.04	0.6699	1.0635	0.1698	4.4461
0.06	0.6612	1.0601	0.1937	4.0082
0.08	0.5956	1.0327	0.2241	3.6467
0.1	0.5782	1.0251	0.2448	3.4069

Table 11. Isothermal compressibility (β_T), Isothermal expansion co-efficient (α), Free volume (V_f) and Internal pressure (π_i) of Pyrazoline derivatives with DMF at 303.15 K.

Concentration	Pyrazoline derivatives (P1-P10)			
	β_T	α	V_f	π_i
(M)	($10^{-14} \text{m}^2 \text{N}^{-1}$)	(10^4K^{-1})	($10^{-6} \text{m}^3/\text{mole}$)	(10^6Pa)
Compound P4				
0.02	0.6634	1.0609	0.1537	4.8820
0.04	0.6430	1.0527	0.1737	4.4479
0.06	0.6364	1.0500	0.1969	4.0265
0.08	0.6047	1.0366	0.2221	3.7024
0.1	0.5841	1.0277	0.2405	3.4755
Compound P5				
0.02	0.6908	1.0717	0.1584	4.8644
0.04	0.6513	1.0561	0.1799	4.4531
0.06	0.6202	1.0432	0.1967	4.1461
0.08	0.5955	1.0327	0.2146	3.8763
0.1	0.5598	1.0168	0.2392	3.6092
Compound P6				
0.02	0.7172	1.0818	0.1524	4.9480
0.04	0.6453	1.0536	0.1770	4.5026
0.06	0.6172	1.0420	0.1941	4.1968
0.08	0.5999	1.0346	0.2113	3.9339
0.1	0.5585	1.0162	0.2311	3.6929

Table 12. Isothermal compressibility (β_T), Isothermal expansion co-efficient (α), Free volume (V_f) and Internal pressure (π_i) of Pyrazoline derivatives with DMF at 303.15 K.

Concentration	Pyrazoline derivatives (P1-P10)			
	β_T	α	V_f	π_i
(M)	($10^{-14} \text{m}^2 \text{N}^{-1}$)	(10^4K^{-1})	($10^{-6} \text{m}^3/\text{mole}$)	(10^6Pa)
Compound P7				
0.02	0.6936	1.0728	0.1483	4.9641
0.04	0.6566	1.0582	0.1696	4.5187
0.06	0.6227	1.0443	0.1888	4.1742
0.08	0.5951	1.0325	0.2046	3.9095
0.1	0.5615	1.0176	0.2256	3.6489
Compound P8				
0.02	0.7123	1.0800	0.1513	4.9375
0.04	0.6968	1.0741	0.1707	4.5165
0.06	0.6639	1.0611	0.1850	4.2181
0.08	0.6548	1.0575	0.2050	3.8891
0.1	0.6290	1.0469	0.225	3.6432

Compound P9				
0.02	0.7290	1.0832	0.1500	4.9601
0.04	0.6919	1.0721	0.1699	4.5256
0.06	0.6547	1.0574	0.1922	4.1547
0.08	0.6254	1.0454	0.2158	3.8292
0.1	0.6091	1.0385	0.2346	3.5754
Compound P10				
0.02	0.6723	1.0645	0.1545	4.9151
0.04	0.6348	1.0493	0.1754	4.4850
0.06	0.6241	1.0449	0.1968	4.1084
0.08	0.6122	1.0399	0.2165	3.8138
0.1	0.5994	1.0344	0.2374	3.5413

4. RESULTS AND DISCUSSION

From the Tables (1-3) it is known that, the ultrasonic velocity (U), density (ρ) and viscosity (η) increases with increase in concentration. Based on the model for sound propagation proposed by Eyring and Kincaid [23], ultrasonic velocity should increase, if the inter molecular free length decreases and vice versa. The linear variation of density and viscosity indicates that there exist a strong interaction between solute and solvent.

In fact, the molecular association increases ultrasonic velocity (u) and acoustic impedance (Z), decreases intermolecular free length (L_f) and adiabatic compressibility (β). A reduction in adiabatic compressibility (K_s) is an indication that component molecules are held close to each other. The decrease in the values of adiabatic compressibility (β) and inter molecular free length (L_f) with increase in ultrasonic velocity (u) further strengthens the strong molecular association between the unlike molecules through hydrogen bonding.

Acoustic impedance increases with increase in concentration. Specific acoustic impedance is directly proportional to ultrasonic velocity and inversely proportional to adiabatic compressibility and shows similar behaviour to that of ultrasonic velocity and opposite to that of adiabatic compressibility [24]. Decrease in adiabatic compressibility might be due to aggregation of solvent molecules around solute molecules. Non linear variation of adiabatic compressibility as a function of composition on of liquid mixture is sufficient evidence for existence of molecular interactions in solutions [25-27].

Intermolecular free length decreases with increase in concentration. The decrease in free length is due to the close packing of the molecules inside the shield, which may be brought by strengthening of molecular interactions. This may be attributed to the fact that the intermolecular interactions might have resulted in a decreased intermolecular free length and a compact structural arrangement.

The Relative association decreases with increase in concentration. Decrease in relative association which indicates the breaking up of the solvent molecules on addition of solute [28]. The relaxation time decreases with increase in concentration. The variations in specific relaxation time are mainly due to the change in viscosity of solutions due to both concentration and temperature. The dispersion of the ultrasonic velocity in the system may contain information about the characteristic time (τ) of the relaxation process that causes dispersion.

Rao's constant and molar compressibility increases with increase in concentration which indicates that the magnitude of interactions is enhanced. The ultrasonic attenuation

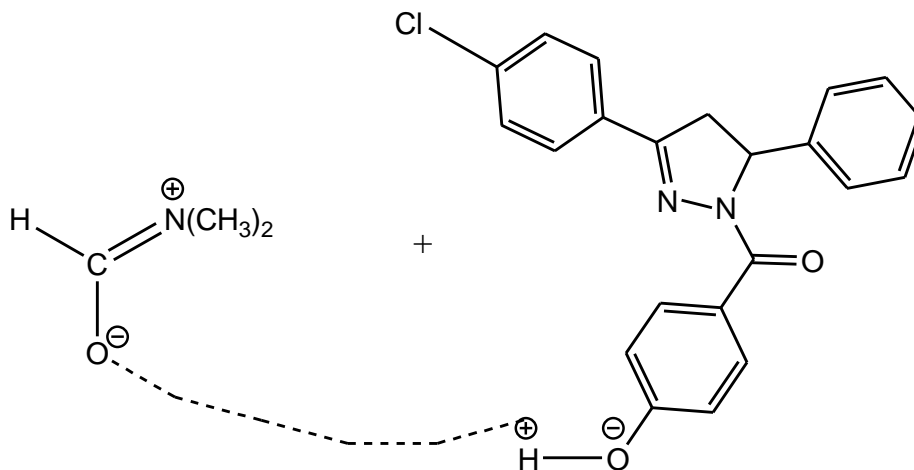
decreases with increase in concentration. The variation of ultrasonic attenuation with increase in concentration is non-linear. This non-linear variation of absorption with concentration of one component strongly supports the presence of strong inter molecular interaction [29]. The van der Waals constant increases with increase in concentration and the Isothermal compressibility decreases with increase in concentration of the solute.

Internal pressure in a liquid system is a measure of intermolecular cohesive forces. The internal pressure decreases with increase in concentration of solute, which indicates the decrease in cohesive forces. This suggests close packing of the molecules inside the shield, which may be brought about by the increasing magnitude of interactions [30-31].

The free volume increases with increase in concentration. The decrease in molecular association of solvent molecule causes an increase in free volume. This may be explained that there is a tendency for the solute molecules to move away from each other, reducing the possibility for interaction, which may further reduce the cohesive force and ultimately lead to an increase in free volume [32].

5. INTERACTIONS IN PYRAZOLINE – DIMETHYLFORMAMIDE

Dimethylformamide is polar in nature. In general, the –OH group is of particular interest because of its highly polar in nature. The associative hydroxyl group in Pyrazoline compound act as proton donor enabling hydrogen bonding with dimethylformamide.



In the above system studied, there is possibility for three types of interactions. (i) Intermolecular hydrogen bonding takes place between carbonyl oxygen of dimethylformamide with hydroxyl hydrogen of pyrazoline compound. (ii) a dipole-dipole interaction takes place between carbonyl carbon of dimethylformamide and hydroxyl oxygen of Pyrazoline compound (iii) an electron donor - electron acceptor complex formation takes between the electron donating methyl group of dimethylformamide and the electron withdrawing chlorine atom of Pyrazoline compound.

So there is existence of solute-solvent interactions between the solvent (dimethylformamide) and solute (pyrazoline compound).

6. CONCLUSION

The acoustical parameters were calculated from the measured ultrasonic velocity, density and viscosity. The existence of molecular interactions in solute-solvent is favoured in the system, which is confirmed by the Complex formations, formation of hydrogen bond, dipole – dipole interactions in solutions and their effect in physical properties of the mixture. Hence it is concluded that there exist a strong molecular interaction between the solute (Pyrazoline derivatives) and the solvent (DMF).

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(Received 03 September 2014; accepted 16 September 2014)